AWARD LIST
SCREENING TEST (WRITTEN)
FOR THE POST OF TRAINED GRADUATE TEACHERS (TGT) (COTNRACT) UNIVERSITY MODEL SCHOOL, UNIVERSITY OF PESHAWAR

Dated.20.01.2023

| S\# | Name of applicant with parentage | Subject | Marks |  |
| :---: | :---: | :---: | :---: | :---: |
| 1. | Durr-e-Shahwar Zafar D/O Mr. Zafar Mumtaz | Chemistry | $30$ | $6$ |
| 2. | Farah Jamshed D/O Qazi Jamshedadam Ali | Chemistry | - |  |
| 3. | Halima Bibi D/O Sher Bahadar Khan | Chemistry | $96$ | $b$ |
| 4. | Huma Iqbal D/O Mr. Muhammad Iqbal | Chemistry | $/ S$ |  |
| 5. | Laila Gulfam D/O Mr. Gulfam | Chemistry | $17$ |  |
| 6. | Maheen Rahim D/O Mr. Fazal e Rahim | Chemistry |  | $4.6$ |
| 7. | Saima Momin D/O Mr. Momin Khan | Chemistry | $23$ | $b \cdot b$ |
| 8. | Sana Nabi D/O Mr. Muhammad Nabi | Chemistry | $33$ | $b-\dot{b}$ |
| 9. | Shaista Zeb D/O Mr. Abdul Saboor Khan | Chemistry | $91$ | $4 \cdot 2$ |
| 10. | Shumaila Javed D/O Mr. Javed Iqbal | Chemistry |  |  |
| 11. | Sumera Rehman D/O Mr. Sher Rehman | Chemistry | $20$ |  |
| 12. | Wagma D/O Mr. Momin Khan | Chemistry | $93$ | $4-6$ |
| 13. | Rozina D/O Mr. Sher Daraz Khan | Chemistry | $29$ | $5-6$ |

## AWARD LIST SCREENING TEST (WRITTEN) FOR THE POST OF TRAINED GRADUATE TEACHERS (TGT) (COTNRACT) UNIVERSITY PUBLIC SCHOOL, UNIVERSITY OF PESHAWAR

Dated.20.01.2023

| S\# | Name of applicant with parentage | Subject | Marks |
| :---: | :--- | :--- | :--- |
| 1. | Bakhtawara D/O Malik Abdul baser Khan | Chemistry |  |
| 2. | Huma Iqbal D/O Mr. Muhammad Iqbal | Chemistry |  |
| 3. | Muhammad Naveed S/O Mr. Jehangir <br> Khan | Chemistry |  |
| 4. | Muhammad Saqib S/O Mr. Muhammad <br> Shahab | Chemistry |  |
| 5. | Qaisar Khan S/O Mr. Ayub Khan | Chemistry |  |
| 6. | Rozina D/O Mr. Sher Daraz Khan | Chemistry |  |
| 7. | Saif Ullah Khan S/O Haji Lal Khan | Chemistry |  |
| 8. | Sana Nabi D/O Mr. Muhammad Nabi | Chemistry |  |



$$
\begin{gathered}
\text { DUR- E SHEHWAR AFAR } \\
D / O A F A R \text { MUMTA2 } \\
17301-8670554-0
\end{gathered}
$$

## Screening test for the appointment of TGT (Contract)

Subject: Chemistry
Note: Choose the correct answer and all questions are compulsory

Q-1. The process of identifying the component present in a sample is called

A Quantitative analysis
(B) Qualitative analysis
C. Volumetric analysis
D. Gravimetric analysis

Q-2. Conductometry is based on
A. Electric current
B. Electrical potential
C. Dielectric constant
$\checkmark$ D. Electrical conductance
Q-3. Which of the following technique is based on the emission of light radiation?
A. Flame photometry
B. Atomic absorption spectrophotometry
C. Ramen spectroscopy
D. Conductometry

Q-4. Which of the following method is based on the solubility difference between the analyte and the unwanted components
A. Distillation
B. Complex formation
C. Electrodeposition
(D) Precipitation

Q-5. Which of the following technique is based on the rate of reaction?
$\sqrt{\text { A. Isotopic dilution analysis }}$
B. Flow injection analysis
C. Mass spectrometry
D. None

Q-6. The term accuracy refers to how near the observed value is to

A) Mean value
B. Significant value
C. Actual value
D. None

Q-7. When HCl is titrated aga inst NaOH the pH at equivalence point will be
A. None
(B) equal to 7
C. less than 7
D. greater than 7

Q-8. When CH 3 COOH is titrated against NaOH the pH at the end point is against NaOH
A. None
(B) equal to 7
C. less than 7
D. greater than 7


Q-9. Which of the following is best indicator for titration of $\mathrm{NH}_{4} \mathrm{OH}$ with HCl ?
A. Methyl red
B. Methyl orange
(C) Phenolphthalein
D. None


Q-10. Soap is a salt of compound known as
A. Acetic acid
B. Formic acid
(C.) Fatty acid
D. Amino acid

Q-11. The colour of water in a lake is due to
A. Excessive growth of sea weeds
(B.) Algae
C. Grass
D. Other pollution

Q-12. Cement sa mixture of
A. Clay and clinker
B. Clay, limestone and gypsum
C. Limestone and gypsum
D. Limestone and clay

Q-13. The percentage of nitrogen in urea is
A. $36 \%$
B. $46 \%$
(C) $56 \%$
D. $66 \%$


Q-14. The percentage of nitrogen in ammonia is
A. $32 \%$


Q-15. Copper occur in nature as
A. Native

B Combined
C. Both
D. None


Q-16. The formula of Cryolite is
A. $\mathrm{Na}_{3} \mathrm{AlF}_{3}$
B. $\mathrm{Na}_{3} \mathrm{AlF}_{4}$
C. $\mathrm{Na}_{3} \mathrm{AlF}_{5}$
D. None

$x$

Q 17. Cell with nuclear
Q-17. Cells with nuclear envelop-belongs to
A. Prokaryotes
(B) Eukaryotes

C. Both
D. None

Q-18. The diameter of a typical plant or animal cell is
A. 1 to $5 \mu \mathrm{~m}$
B. 5 to $10 \mu \mathrm{~m}$
C. 10 to $50 \mu \mathrm{~m}$
D. 5 to $100 \mu \mathrm{~m}$


Q-19. Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by
A. Davy
(B) Liebig
C. Boyle
D. Rouelle


Q-20. Which of the following statement is not true regarding buffers?
A. They are made up of weak acid and its salts
B. They are made up of weak base and its salts
C. they tend to resist a change in pH value
D. they are made up of strong acid and weak base

Q-21. Relative order of acidity of $\mathrm{HF}, \mathrm{HCl}, \mathrm{HBr}$ and HI acids is
A. $\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}>\mathrm{HF}$
B. $\mathrm{HF}>\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}$
(C. $\mathrm{HI}>\mathrm{HBr}>\mathrm{HCl}>\mathrm{HF}$
D. $\mathrm{HF}>\mathrm{Hi}>\mathrm{HCl}>\mathrm{HBr}$

Q-22. The pH of 0.1 M HCl solution is $\begin{array}{llll}\text { A. } 0.01 & \text { B. } 0.1 & \text { C. } 0.2 & \text { (D.) } 1\end{array}$

Q-23. Which of the following amino acid have two carboxylic groups?
(A. Aspartic acid
B. Histidine
C. Gluatmine
D. none

Q-24. How many isomeric aldoses are possible for the molecular formula $\mathrm{C}_{6} \mathrm{H}_{12} \mathrm{O}_{6}$
(A) 2
B. 4
C. 6
D. 16

Q-25. Epimers are compounds that differ in
A. functional group
(8.) configuration at alpha carbon
C. configuration at any carbon
D. none

Q-26. Select a basic amino acid
A. Glycine
B. Cystine
(C) Alanine
D. Lysine


Q-27. Apoenzyme is
A. Hydrolytic enzyme
B. Oxidative enzyme
C. Coenzyme
(D) Protein part of enzyme after removal of coenzyme

Q-28. A sugar present in DNA is
A. D-ribose
B. D-glucose
(C) 2-Deoxy-D-Ribose
D. None

Q-29. Which vitamin is responsible for vision?
(A) Retinol
B. Vit D
C. Tocopherols
D. Vit. K

Q-30. Which of the following is periodic propery?
A. Atomic volume
B. Metallic character
C. Ionization energy
(D) All

Q-31. Which of the following has largest size
A. $\mathrm{Na}^{+}$
(B.) $\mathrm{Cl}^{-}$
C. $F^{-}$
D. $\mathrm{Fe}^{2+}$

Q-32. Which group contains elements that exist as monoatomic molecules?
(D) 18
r
B. 2 C. 14
A. 1
$\qquad$
Q-33. Which of the following element is more metallic?
B. As
C. Sb
(D. Bi

Q-34. Which of the following factor effect ionization energy?
A. Atomic radius
C. Shiedding effect
B. Nuclear charge

Q-35. Which of the following element is metalloid?
A. Pb
B. carbon C. As
D. Mg


Q-36. Which of the following compound does not following octet rule?
B. $\mathrm{PBr}_{3}$
C. IBr
(D) $\mathrm{BrF}_{5}$

Q-37. Which of the following halides have the
lowest melting point?
$\begin{array}{ll}\text { A. } \mathrm{NaCl} & \mathrm{B} . \mathrm{NaF}\end{array}$
C. NaBr
(D) NaI

Q-38. Which of the following halides has zero
dipole moment? dipole moment?
A. $\mathrm{NH}_{3}$
(B) $\mathrm{Cl}_{2}$
C. $\mathrm{NF}_{3}$
D. $\mathrm{CHCl}_{3}$

Q-39. The state of hybridization of carbon in CO 2 is?
A Sp
(A) sp
B. sp 2
C. sp 3
D. dsp 2

Q-40. Which one of the following is non-polar?
A. $\mathrm{CH}_{2} \mathrm{Cl}$
(B. $\mathrm{CCl}_{4}$
C. $\mathrm{CHCl}_{3}$
D. $\mathrm{CHCl}_{3}$


Q-41. Silvite is an ore of..?
A. Ca
B. Mg
C. Ba
(D) K

Q-42. Which elements are non metals?
A. $N \& P \quad B . A s \& S b$
C. $\mathrm{Sb} \& \mathrm{Bi}$
D. $\mathrm{Ba} \& \mathrm{Bi}$

Q-43. Among oxides of nitrogen all are gases except
(A) N 2 O 5
B. N 2 O
C. NO
D. N 2 O 3

Q-44. All halogens form oxyacid except
A. Flourine B. Chlorine C. Bromine ( 0 Iodine
Q-45. Compounds HCN and HNC are
A. Tautomer
C. functional isomers
(B.) Metamer
D. Conformers


## Name $\rightarrow$ Laila Gulfam

## Screening test for the appointment of TGT (Contract)

 Subject: ChemistryNote: Choose the correct answer and all questions are compulsory
,
Q-1. The process of identifying the component present in a sample is called

LA. Quantitative analysis
B. Qualitative analysis
C. Volumetric analysis
D. Gravimetric analysis

Q-2. Conductometry is based on
A. Electric current
B. Electrical potential
C. Dielectric constant
D. Electrical conductance

Q-3. Which of the following technique is based on the emission of light radiation?
A. Flame photometry

LB. Atomic absorption spectrophotometry
C. Ramen spectroscopy
D. Conductometry

Q-4. Which of the following method is based on the solubility difference between the analyte and the unwanted components
A. Distillation
B. Complex formation
C. Electrodeposition
D.. Precipitation

Q-5. Which of the following technique is based on the rate of reaction?
A. Isotopic dilution analysis
B. Flow injection analysis
C. Mass spectrometry
D. None

Q-6. The term accuracy refers to how near the observed value is to
A. Acetic acid
B. Formic acid

1. Fatty acid
D. Amino acid

Q-11. The colour of water in a lake is due to
A. Excessive growth of sea weeds
V. Algae
C. Grass
D. Other pollution

Q-12. Cement sa mixture of
A. Clay and clinker

LB. Clay, limestone and gypsum
C. Limestone and gypsum
D. Limestone and clay

Q-13. The percentage of nitrogen in urea is
B. $46 \%$
C. $56 \%$
D. $66 \%$

Q-14. The percentage of nitrogen in ammonia is
A. $32 \%$
B. $42 \%$
e. $72 \%$
D. $82 \%$


Q-15. Copper occur in nature as
A. Native
B. Combined
C. Both
D. None


Q-16. The formula of Cryolite is
A. $\mathrm{Na}_{3} \mathrm{AlF}_{3}$

LB. $\mathrm{Na}_{3} \mathrm{AlF}_{4}$
C. $\mathrm{Na}_{3} \mathrm{AlF}_{5}$
D. None

Q-17. Cells with nuclear envelop belongs to
A. Prokaryotes
V. Eukaryotes
C. Both
D. None

Q-18. The diameter of a typical plant or animal cell is


LA. 1 to $5 \mu \mathrm{~m}$
B. 5 to $10 \mu \mathrm{~m}$
C. 10 to $50 \mu \mathrm{~m}$
D. 5 to $100 \mu \mathrm{~m}$

Q-19. Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by

A. Davy
B. Liebig
C. Boyle
D. Rouelle

Q-20. Which of the following statement is not true regarding buffers?

A. They are made up of weak acid and its salts B. They are made up of weak base and its salts . they tend to resist a change in pH value
D. they are made up of strong acid and weak base

Q-21. Relative order of acidity of $\mathrm{HF}, \mathrm{HCl}, \mathrm{HBr}$ and Hl acids is
A. $\mathrm{HCl}>\mathrm{HBr}>\mathrm{Hl}>\mathrm{HF}$

LB. $\mathrm{HF}>\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}$
C. $\mathrm{Hl}>\mathrm{HBr}>\mathrm{HCl}>\mathrm{HF}$
D. $\mathrm{HF}>\mathrm{HI}>\mathrm{HCl}>\mathrm{HBr}$
$V$
Q-22. The pH of 0.1 M HCl solution is
Q. 0.01
B. 0.1
C. 0.2
D. 1

Q-23. Which of the following amino acid have two carboxylic groups?

A. Aspartic acid

LB. Histidine
C. Gluatmine
D. none

Q-24. How many isomeric aldoses are possible for the molecular formula $\mathrm{C}_{6} \mathrm{H}_{12} \mathrm{O}_{6}$
LA. 2
B. 4
C. 6
D. 16

Q-25. Epimers are compounds that differ in

- A. functional group
B. configuration at alpha carbon
C. configuration at any carbon
D. none

Q-26. Select a basic amino acid
A. Glycine
B. Cystine
C. Alanine
D. Lysine

Q-27. Apoenzyme is
A. Hydrolytic enzyme
B. Oxidative enzyme
C. Coenzyme
D. Protein part of enzyme after removal of coenzyme

Q-28. A sugar present in DNA is
A. D-ribose
B. D-glucose
C. 2-Deoxy-D-Ribose
D. None
Q-29. Which vitamin is responsible for vision?
A. Retinol
B. Nit $D$
C. Tocopherols
D. Vic. K

Q -30 . Which of the following is periodic propery?
A. Atomic volume
B. Metallic character
C. Ionization energy
D. All

Q-31. Which of the following has largest size ?
A. $\mathrm{Na}^{+}$
B. $\mathrm{Cl}^{-}$
LC
D. $\mathrm{Fe}^{2+}$

Q-32. Which group contains elements that exist as monoatomic molecules?
B. 2
C. 14
D. 18

Q-33. Which of the following element is more metallic?
A. P
B. As
c. Sb
D. Bi

Q-34. Which of the following factor effect ionization energy?
A. Atomic radius
B. Nuclear charge
C. Shiedding effect


Q-35. Which of the following element is metalloid?
A. Pb
B. carbon
C. As
D. Mg
$N$
Q-36. Which of the following compound does not
following octet rule?
A. $\mathrm{CS}_{2}$
B. $\mathrm{PBr}_{3}$
b. IBr
D. $\mathrm{BrF}_{5}$

Q -37 . Which of the following halides have the lowest melting point?
A. NaCl
B. NaF
C. NaBr
D. Na

X Q-38. Which of the following halides has zero dipole moment?
A. $\mathrm{NH}_{3}$
B. $\mathrm{Cl}_{2}$
C. $\mathrm{NF}_{3}$
D. $\mathrm{CHCl}_{3}$

Q-39. The state of hybridization of carbon in CO 2 is?
B. sp 2
C. sp 3
D. dsp 2

Q-40. Which one of the following is non-polar?
A. $\mathrm{CH}_{2} \mathrm{Cl}$
B. $\mathrm{CCl}_{4}$
C. $\mathrm{CHCl}_{3}$
D. $\mathrm{CHCl}_{3}$

K
Q-41. Silvite is an ore of..?
A. Ca
B. Mg
C. Ba


Q-42. Which elements are non metals? $\begin{array}{llll}L A . N \& P & \text { B. As \& } S b & \text { C. } S b \& B i & D . B a \& B i\end{array}$

Q-43. Among oxides of nitrogen all are gases except
A. N2O5
B. N 2 O
C. NO


Q-44. All halogens form oxyacid except
A. Flouring B. Chlorine $C$. Bromine D. Iodine

Q-45. Compunds HCN and HNC are
A. Tautomer
B. Metamer
16. Functional isomers
D. Conformers
Q-46. Cetane is
A. n-hexane
VB. n-pentadecane
C. n-octane
D. n-hexadecane

Q-47. Which among the following is not nucleophile?
A. CH 3 NH 3
B. $\mathrm{CH}_{2} \mathrm{CH}_{2}$
C. $\mathrm{CH}_{3}+$
D. OH

Q-48. Which of the following is strong acid?
A. Benzoic acid
B. m-nitrobenzoic acid
C. o-nitrobenzoic acid
D. p-nitrobenzoic acid

Q-49. Which of the following orbital does not make sense?
A. $6 f$
B. 4 f
C. 7 s


Q-50. The maxium value of order of reaction may be?
A. 1

B. 2
D. 4

3

## Name Sana Nab,

 Father Name. Muhammad Nab.: S.NO

## Screening test for the appointment of TGT (Contract)

Subject: Chemistry
Note: Choose the correct answer and all questions are compulsory.

Q-1. The process of identifying the component present in a sample is called
A. Quantitative analysis
18. Qualitative analysis
C. Volumetric analysis
D. Gravimetric analysis

Q-2. Conductometry is based on
A. Electric current
B. Electrical potential
C. Dielectric constant
V. Electrical conductance

Q-3. Which of the following technique is based $\square$ the emission of light radiation?
A. Flame photometry

B Atomic absorption spectrophotometry
C. Raman spectroscopy
D. Conductometry
Q.4. Which of the following method is based on the solubility difference between the analyse and the unwanted components
A. Distillation
B. Complex formation
C. Electrodeposition
D. Precipitation

Q-5. Which of the following technique is based on the fate of reaction?

L6. Flow injection, analysis
C. Mass spectrometry
D. None

Q-6. The term accuracy refers to how near the observed value is to
A. Mean value
B. Significant value
C. Actual value
D. None

Q-7. When HCl is titrated against NaOH the pH
at equivalence point will be
A. None
B. equal to 7
C. less than 7
D. greater than 7

Q-8. When CH 3 COOH is titrated against NaOH the pH at the end point is
A. None
B. equal to 7
C. less than 7
$\sqrt{ }$. greater than 7
Q-9. Which of the following is best indicator for titration of $\mathrm{NH}_{4} \mathrm{OH}$ with HCl ?
A. Methyl red
B. Methyl orange
C. Phenolphthalein
D. None

Q-10. Soap is a salt of compound brown as
A. Acetic acid
E. Formic acid
2. Fatty acid
D. Amino acid
Q.11. The colour of water in a lake is due to

LA. Excessive growth of sea weeds
B. Algae
C. Grass
D. Other pollution

Q-12. Cements a mixture x . ,

B. Clay, limestone and gypsum
C. Limestone and gypsum
D. Limestone and clay

Q-13. The percentage of nitrogen in urea is
A. $36 \%$
B. $46 \%$
C. $56 \%$
C. $56 \%$

- $66 \%$

Q-74. The percentage of mitiogen in ammon is
A. $32 \%$
B. $42 \%$
C. $72 \%$
b. $82 \%$

Q-15. Copper occur in nature as
A. Native
B. Combined
C. Both
D. None


Q-16. The formula of Cryolite is
VA. $\mathrm{Na}_{3} \mathrm{AlF}_{3}$

B. $\mathrm{Na}_{3} \mathrm{AlF}_{4}$
C. $\mathrm{Na}_{3} \mathrm{AlF}_{5}$
D. None

Q-17. Cells with nuclear envetop belongs to
A. Prokaryotes
B. Eukaryotes
C. Both
D. None

Q-18. The diameter of a typical plant or animal cell is


Q-19. Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by
A. Davy
B. Liebig
C. Boyle
D. Rouelle


Q-20. Which of the following statement is not true regarding buffers?
A. They are made up of weak acid and its salts
B. They are made up of weak base and its salts
C. they tend to resist a change in pH value

Dase they are made up of strong acid and weak
Q-21. Relative order of acidity of $\mathrm{HF}, \mathrm{HCl}, \mathrm{HBr}$ and HI acids is
A. $\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}>\mathrm{HF}$
B. $\mathrm{HF}>\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}$
e. $\mathrm{HI}>\mathrm{HBr}>\mathrm{HCl}>\mathrm{HF}$
D. $\mathrm{HF}>\mathrm{HI}>\mathrm{HCl}>\mathrm{HBr}$


Q-22. The pH of 0.1 M HCl solution is
A. 0.01
B. 0.1
C. 0.2
D. 1

Q-23. Which of the following amino acid have two carboxylic groups?
*. Aspartic acid

C. Gluatmine
D. none

Q-24. How many isomeric aldoses are possible for the molecular formula $\mathrm{C}_{6} \mathrm{H}_{12} \mathrm{O}_{6}$
A. 2
B. 4
C. 6
b. 16


Q-25. Epimers are compounds that differ in
A. functional group
V. configuration at alpha carbon
C. configuration at any carbon
D. none

Q-26. Select a basic amino acid
A. Glycine
B. Cystine C. Alanine
D. Lysine


Q-27. Apoenzyme is
A. Hydrolytic enzyme
B. Oxidative enzyme
V. Coenzyme
D. Protein part of enzyme after removal of coenzyme

Q-28. A sugar present in DNA is
A. D-ribose
B. D-glucose
V. 2-Deoxy-D-Ribose
D. None

Q-29. Which vitamin is responsible for vision?
A. Retinol
B. Vit D
C. Tocopherols
D. Vit. K


Q-30. Which of the following is periodic propery?
A. Atomic volume
C. Ionization energy
B. Metallic character


Q-31. Which of the following has largest size ?
A. $\mathrm{Na}^{+}$
B. $\mathrm{Cl}^{-}$
C. $\mathrm{F}^{-}$
D. $\mathrm{Fe}^{2+}$

Q-32. Which group contains elements that exist as monoatomic molecules?
A. 1
B. 2
C. 14


Q-33. Which of the following element is more metallic?
A. P
B. As
C. Sb
2. Bi

Q-34. Which of the following factor effect ionization energy?
A. Atomic radius
B. Nuclear charge
C. Shiedding effect
D. Ail

Q35. Which of the following element is metalloid?
A. Pb
B. carbon
c. As
D. Mg

Q36. Which of the following compound does not following octet rule?
A. $\mathrm{CS}_{2}$
B. $\mathrm{PBr}_{3}$
C. IBr
V. $\mathrm{BrF}_{5}$

Q-37. Which of the following halides have the lowest melting point?
A. NaCl
B. NaF
C. NaBr
D. NaI

Q-38. Which of the following halides has zero dipole moment?
A. $\mathrm{NH}_{3}$
C. $\mathrm{NF}_{3}$
D. $\mathrm{CHCl}_{3}$
L. $\mathrm{Cl}_{2}$

Q-39. The state of hybridization of carbon in CO 2 is?
A. sp
B. sp 2
C. sp 3
D. dsp 2

Q40. Which one of the following is non-polar?
A. $\mathrm{CH}_{2} \mathrm{Cl}$ B. $\mathrm{CCl}_{4}$
C. $\mathrm{CHCl}_{3}$
D. $\mathrm{CHCl}_{3}$

Q-41. Silvite is an ore of..?
A. Ca
V. Mg
C. Ba
D. $K$

Q-42. Which elements are non metals?
$\begin{array}{llll}\text { A. } N \& P & \text { B. As \& } \mathrm{Sb} & \text { C. } \mathrm{Sb} \text { \& } \mathrm{Bi} & \mathrm{D} . \mathrm{Ba} \& \mathrm{Bi}\end{array}$
Q-43. Among oxides of nitrogen all are gases
except
R. N2O5
B. N 2 O
C. NO
D. N 2 O 3

Q-44. All halogens form oxyacid except
A. Flouring
B. Chlorine
C. Bromine V. Iodine

Q-45. Compounds HCN and HNC are
A. Tautomers
B. Metamer

C functional isomers
D. Conformers

$$
\text { Name }=\text { Shasta, Web } D / 0
$$



Screening test for the appointment of TGT (Contract)
Subject: Chemistry
Note: Choose the correct answer and all questions are compulsory.

Q-1. The process of identifying the component present in a sample is called
A. Quantitative analysis
(B) Qualitative analysis
C. Volumetric analysis
D. Gravimetric analysis

Q-2. Conductometry is based on
A. Electric current
B. Electrical potential
e. Dielectric constant
D) Electrical conductance

Q-3. Which of the following technique is based on the emission of light radiation?
A. Flame photometry
(B) Atomic absorption spectrophotometry
C. Ramen spectroscopy
D. Conductometry

Q-4. Which of the following method is based on the solubility difference between the analyte and the unwanted components
A. Distillation
B. Complex formation
C. Electrodeposition
(D) Precipitation

Q-5. Which of the following technique is based on the rate of reaction?
A. Isotopic dilution analysis
B. Flow injection analysis
C. Mass spectrometry
(D) None

Q-6. The term accuracy refers to how near the observed value is to
A. Mean value
B. Significant value
C. Actual value
D. None

Q-7. When HCl is titrated against NaOH the pH f at equivalence point with be
A. None
B. equal to 7
(C) less than 7
D. greater than 7

Q-8. When CH 3 COOH is titrated against NaOH the pH at the end point is
A. None
B. equal to 7
C. less than 7
(D) greater than 7

Q-9. Which of the following is best indicator for titration of $\mathrm{NH}_{4} \mathrm{OH}$ with HCl ?
A. Methyl red
B. Methyl orange
(C) Phenolphthaleine
D. None

Q-10. Soap is a salt of compound known as
A. Acetic acid
B. Formic acid
(C) Fatty acid
D. Amino acid

Q-11. The colour of water in a lake is due to
A. Excessive growth of sea weeds
(B) Algae
C. Grass
D. Other pollution

Q-12. Cement sa mixture of
A. Clay and clinker
(B). Clay, limestone and gypsum L
C. Limestone and gypsum
D. Limestone and clay

Q-13. The percentage of nitrogen in urea is
A. $36 \%$
B. $46 \%$
C. $56 \%$
D. $66 \%$


Q-14. The percentage of nitrogen in ammonia is
A. $32 \%$
(B) $42 \%$
C. $72 \%$
D. $82 \%$

Q-15. Copper occur in nature as
A. Native
(B) Combined
C. Both
D. None


Q-16. The formula of Cryolite is
(A.) $\mathrm{Na}_{3} \mathrm{AlF}_{3}$
B. $\mathrm{Na}_{3} \mathrm{AlF}_{4}$
C. $\mathrm{Na}_{3} \mathrm{AlF}_{5}$
D. None

Q-17. Cells with nuclear envelop belongs to
A. Prokaryotes
B. Eukaryotes
C. Both
D. None

Q-18. The diameter of a typical plant or animal cell is
(A.) 1 to $5 \mu \mathrm{~m}$
B. 5 to $10 \mu \mathrm{~m}$
C. 10 to $50 \mu \mathrm{~m}$
D. 5 to $100 \mu \mathrm{~m}$


Q-19. Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by
A. Davy
(B) Liebig
C. Boyle

D. Rouelle

Q-20. Which of the following statement is not true regarding buffers?
A. They are made up of weak acid and its salts
B. They are made up of weak base and its salts
C. they tend to resist a change in pH value
(D) they are made up of tstrong acid and weak base

Q-21. Relative order of acidity of $\mathrm{HF}, \mathrm{HCl}, \mathrm{HBr}$ and HI acids is
A. $\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}>\mathrm{HF}$
B. $\mathrm{HF}>\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}$
(C) $\mathrm{Hl}>\mathrm{HBr}>\mathrm{HCl}>\mathrm{HF}$
D. $\mathrm{HF}>\mathrm{HI}>\mathrm{HCl}>\mathrm{HBr}$

Q-22. The pH of 0.1 M HCl solution is
A. 0.01 (B) 0.1
C. 0.2
D. 1


Q-23. Which of the following amino acid have two carboxylic groups?
A. Aspartic acid
B) Histidine
C. Gluatmine
D. none

Q-24. How many isomeric aldoses are possible for the molecular formula $\mathrm{C}_{6} \mathrm{H}_{12} \mathrm{O}_{6}$
(A) 2
B. 4
C. 6
D. 16


Q-25. Epimers are compounds that differ in
(A.) functional group
B. configuration at alpho carbon
C. configuration at any/carbon
D. none

Q-26. Select a basic amino acid
A. Glycine
B. Cystine
(C) Alanine
D. Lysine


Q-27. Apoenzyme is
A. Hydrolytic enzyme
B. Oxidative enzyme
C. Coenzyme

(D.) Protein part of enzyme after removal of coenzyme

Q-28. A sugar present in DNA is
A. D-ribose
B. D-glucose

(C) 2-Deoxy-D-Ribose
D. None

Q-29. Which vitamin is responsible for vision?
(A) Retinol
B. Vit D
C. Tocopherols
D. Vit. K

Q-30. Which of the following is periodic propery?
A. Atomic volume
B. Metallic character
(C) Ionization energy
D. All


Q-31. Which of the following has largest size?
A. $\mathrm{Na}^{+}$
(B) $\mathrm{Cl}^{-}$
C. $F^{-}$
D. $\mathrm{Fe}^{2+}$


Q-32. Which group contains elements that exist as monoatomic molecules?
A. 1
B. 2
C. 14
D. 18


Q-33. Which of the following element is more metallic?
A. P
(B) As
C. Sb
D. Bi



Q-34. Which of the following factor effect ionization energy?
A. Atomic radius
B. Nuclear charge
C. Shedding effect
(D) All

Q-35. Which of the following element is metalloid?
A. Pb
B. carbon
C. As
D. Mg


Q-36. Which of the following compound does not following octet rule?
A. $\mathrm{CS}_{2}$
B. $\mathrm{PBr}_{3}$
C. IBr
D. $\mathrm{BrF}_{5}$

Q-37. Which of the following halides have the lowest melting point?
A. NaCl
B. NaF
C. NaBr
(D) NaI

Q-38. Which of the following halides has zero dipole moment?
A. $\mathrm{NH}_{3}$
(B) $\mathrm{Cl}_{2}$
C. $\mathrm{NF}_{3}$
D. $\mathrm{CHCl}_{3}$


Q-39. The state of hybridization of carbon in CO 2 is?
A. sp
(B) sp 2
C. sp 3
D. dsp 2


Q-40. Which one of the following is non-polar?
A. $\mathrm{CH}_{2} \mathrm{Cl}$
B. $\mathrm{CCl}_{4}$
C. $\mathrm{CHCl}_{3}$
D. $\mathrm{CHCl}_{3}$

Q-41. Silvite is an ore of.?
A. Ca
(B) Mg
C. Ba
D. $K$

Q-42. Which elements are non metals?
(A) $N \& P$
B. As \& Sb
C. $\mathrm{Sb} \& \mathrm{Bi}$
D. $\mathrm{Ba} \& \mathrm{Bi}$


Q-43. Among oxides of nitrogen all are gases
except
(A) N 2 O 5
B. N 2 O
C. NO
D. N 2 O 3

Q-44. All halogens form oxyacid except A. Flouring B. Chlorine ©. Bromine D. Iodine


Q-45. Compounds HCN and HNC are
A. Tautomers
(B) Metamer
C. functional isomers
D. Conformers


## Screening test for the appointment of TGT (Contract)

Subject: Chemistry
Note: Choose the correct answer and all questions are compulsory.

Q-1. The process of identifying the component present in a sample is called
A. Quantitative analysis
B. Qualitative analysis

Le. Volumetric analysis
D. Gravimetric analysis

Q-2. Conductometry is based on
A. Electric current
B. Electrical potential
C. Dielectric constant

LD. Electrical conductance
Q-3. Which of the following technique is based on the emission of light radiation?
A. Flame photometry

IB. Atomic absorption spectrophotometry
C. Raman spectroscopy
D. Conductometry

Q-4. Which of the following method is based on the solubility difference between the analyte and the unwanted components

A. Distillation
B. Complex formation
L.Electrodeposition
D. Precipitation

Q-5. Which of the following technique is based on the rate of reaction?

A. Isotopic dilution analysis
B. Flow injection analysis
e. Mass spectrometry
D. None

Q-6. The term accuracy refers to how near the observed value is to
A. Mean value
B. Significant value
C. Actual value
D. None


Q-7. When HCl is titrated against NaOH they pH at equivalence point will be
A. None
B. equal to 7
C. less than 7
D. greater than 7

Q-8. When CH 3 COOH is titrated against NaOH the pH at the end point is
A. None
B. equal to 7
C. less than 7

LD. greater than 7
Q-9. Which of the following is best indicator for titration of $\mathrm{NH}_{4} \mathrm{OH}$ with HCl ?
A. Methyl red
B. Methyl orange

Le. Phenolphthaleine
D. None

Q-10. Soap is a salt of compound known as
A. Acetic acid

LB. Formic acid
C. Fatty acid
D. Amino acid


Q-11. The colour of water in a lake is due to
A. Excessive growth of sea weeds
1.B. Algae
C. Grass
D. Other pollution

Q-12. Cement s a mixture of
A. Clay and clinker
B. Clay, limestone and gypsum

Le. Limestone and gypsum
D. Limestone and clay

Q-13. The percentage of nitrogen in urea is
A. $36 \%$
-B. $46 \%$
C. $56 \%$
D. $66 \%$

Q-14. The percentage of nitrogen in ammonia is
A. $32 \%$
B. $42 \%$
C. $72 \%$
D. $82 \%$


Q-15. Copper occur in nature as
A. Native

LB. Combined
C. Both
D. None

Q-16. The formula of Cryolite is
La. $\mathrm{Na}_{3} \mathrm{AlF}_{3}$

B. $\mathrm{Na}_{3} \mathrm{AlF}_{4}$
C. $\mathrm{Na}_{3} \mathrm{AlF}_{5}$
D. None

Q-17. Cells with nuclear envelop belongs to
A. Prokaryotes
-B. Eukaryotes

C. Both
D. None

Q-18. The diameter of a typical plant or animal cell is
A. 1 to $5 \mu \mathrm{~m}$
B. 5 to $10 \mu \mathrm{~m}$
C. 10 to $50 \mu \mathrm{~m}$

D. 5 to $100 \mu \mathrm{~m}$

Q-19. Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by
A. Davy
B. Liebig
C. Boyle
D. Rouelle

Q-20. Which of the following statement is not true regarding buffers?
A. They are made up of weak acid and its salts
B. They are made up of weak base and its salts

C they tend to resist a change in pH value D. they are made up of strong acid and weak base

Q-21. Relative order of acidity of $\mathrm{HF}, \mathrm{HCl}, \mathrm{HBr}$ and HI acids is
A. $\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}>\mathrm{HF}$
B. $\mathrm{HF}>\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}$
C. $\mathrm{Hl}>\mathrm{HBr}>\mathrm{HCl}>\mathrm{HF}$
D. $\mathrm{HF}>\mathrm{HI}>\mathrm{HCl}>\mathrm{HBr}$


Q-22. The pH of 0.1 M HCl solution is
A. 0.01
B. 0.1
C. $0.2 \quad$ LD. 1


Q-23. Which of the following amino acid have two carboxylic groups?


Q-24. How many isomeric aldoses are possible for the molecular formula $\mathrm{C}_{6} \mathrm{H}_{12} \mathrm{O}_{6}$
A. 2
$-B .4$
C. 6
D. 16


Q-25. Epimers are compounds that differ in
LA. functional group
B. configuration at alpha carbon
C. configuration at any carbon
D. none

Q-26. Select a basic amino acid
A. Glycine

1. Cystine
C. Alanine
D. Lysine


Q-27. Apoenzyme is
A. Hydrolytic enzyme
B. Oxidative enzyme
U. Coenzyme

D. Protein part of enzyme after removal of coenzyme

Q-28. A sugar present in DNA is

D. None

Q-29. Which vitamin is responsible for vision?
A. Retinol
B. Kit D
C. Tocopherols
D. Vil. K

Q-30. Which of the following is periodic propery?
A. Atomic volume
B. Metallic character
e. Ionization energy
D. All




Q-31. Which of the following has largest size ?
A. $\mathrm{Na}^{+}$
(B. $\mathrm{Cl}^{-}$
C. $F^{-}$
D. $\mathrm{Fe}^{2+}$

Q-46. Cetane is
A. n-hexane
L. n-pentadecane
C. n-octane
D. n-hexadecane


Q-47. Which among the following is not nucleophile?
A. CH 3 NH 3
B. $\mathrm{CH}_{2} \mathrm{CH}_{2}$
C. $\mathrm{CH}_{3}+$


Q-48. Which of the following is strong acid?
A. Benzoic acid
LB. m-nitrobenzoic acid
C. o-nitrobenzoic acid
D. p-nitrobenzoic acid


Q-49. Which of the following orbital does not make sense?
A. $6 f$
B. $4 f$
C. 7 s

- 5 g

Q-50. The maxium value of order of reaction may be?
A. 1
LB. 2
C. 3
D. 4

Q-36. Which of the following compound does not following octet rule?
A. $\mathrm{CS}_{2}$
B. $\mathrm{PBr}_{3}$ Le. TBr
D. $\mathrm{BrF}_{5}$

Q-37. Which of the following halides have the
lowest melting point?
A. NaCl
B. NaF \&. NaBr
D. NaI

Q-38. Which of the following halides has zero
dipole moment?
A. $\mathrm{NH}_{3}$
B. $\mathrm{Cl}_{2}$
C. $\mathrm{NF}_{3}$
D. $\mathrm{CHCl}_{3}$


Q-39. The state of hybridization of carbon in CO 2 is?
A. sp
B. sp 2
C. sp 3
D. dsp 2

Q-40. Which one of the following is non-polar?
A. $\mathrm{CH}_{2} \mathrm{Cl}$
$\mathrm{B}, \mathrm{CCl}_{4}$
C. $\mathrm{CHCl}_{3}$
D. $\mathrm{CHCl}_{3}$


Q-41. Silvite is an ore of..?
A. Ca
B. Mg
16. Ba
D. K
$>$

Q-42. Which elements are non metals?
A. N \& P
B. As \& Sb
C. $\mathrm{Sb} \& \mathrm{Bi}$
D. $\mathrm{Ba} \& \mathrm{Bi}$ X

Q-43. Among oxides of nitrogen all are gases
except
A. N 2 O 5
B. N 2 O
c. NO
D. N 2 O 3


Q-44. All halogens form oxyacid except
$A$. Flouring : $B$. Chlorine C. Bromine D. Iodine $\overline{0}$
d HNC are
Q-45. Compounds HCN and HNC are
A. Tautomers
mers
C. functional isomers
D. Conformers


Halima
$17801-0851387.6$

## Screening test for the appointment of TGT (Contract) <br> Subject: Chemistry

Note: Choose the correct answer and all questions are compulsory.

Q-1. The process of identifying the component present in a sample is called
A) Quantitative analysis
B. Qualitative analysis
C. Volumetric analysis
D. Gravimetric analysis

## Q-2. Conductometry is based on

A. Electric current
B. Electrical potential
C. Dielectric constant
D) Electrical conductance

Q-3. Which of the following technique is based on the emission of light radiation?
A. Flame photometry

B Atomic absorption spectrophotometry
C Daman spectroscopy
D Conductometry
Q-4. Which of the following method is based on the solubility difference between the analyte and the unwanted components
A. Distillation
B. Complex formation
C. Electrodeposition
D) Precipitation

Q-5. Which of the following technique is based on the rate of reaction?

A isotopic dilution analysis
B. Flow injection analysis
C. Mass spectrometry

D None
Q-6. The term accuracy refers to how near the observed value is to
A. Mean value
B. Significant value
C) Actual value

D None

Q-7. When HCl is titrated a inst aOOH the pH at equivalence point withe
A. None
(B) equal to 7
C. less than 7
D. greater than 7

Q-8. When CH 3 COOH is titrated against NaOH the pH at the end point is $\qquad$
A. None
$\vee B$ equal to 7
C less than 7
D. greater than 7

Q-9. Which of the following is best indicator for titration of $\mathrm{NH}_{4} \mathrm{OH}$ with HCl ?
A. Methyl red
B. Methyl orange
C. Phenolphthaleine
D. None

Q-10. Soap is a salt of compound known as
A Acetic acid
B. Formic acid
(C) Fatty acid
D. Amino acid

Q-11. The colour of water in a lake is due to
A. Excessive growth of sea weeds
B. Algae
C. Grass
(D) Other pollution

Q-12. Cement s a mixture of
A. Clay and clinker
B. Clay, limestone and gypsum
C. Limestone and gypsum
D. Limestone and clay

Q-13. The percentage of nitrogen in urea is
A. $36 \%$
B. $46 \%$
C. $56 \%$
D. $66 \%$

Q-14. The percentage of nitrogen in ammonia is
A. $32 \%$
B. $42 \%$
C. $72 \%$

D $82 \%$

Q-15. Copper occur in nature as
A Native
(B) Combined
C Both
D. None


Q-16. The formula of Cryolite is
A. $\mathrm{Na}_{3} \mathrm{AlF}_{3}$
B. $\mathrm{Na}_{2} \mathrm{AlF}_{4}$
C. $\mathrm{Na}_{3} \mathrm{AlF}_{5}$
(D) None


Q-17. Cells with nuclear envelop belongs to
A) Prokaryotes
B Eukaryotes
C. Both
D. None

Q-18. The diameter of a typical plant or animal cell is
A. 1 to $5 \mu \mathrm{~m}$
B. 5 to $10 \mu \mathrm{~m}$

C 10 to $50 \mu \mathrm{~m}$
D. 5 to $100 \mu \mathrm{~m}$


Q-19. Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by
A. Davy
B) Liebig
C. Boyle
D. Rouelle


Q-20. Which of the following statement is not true regarding buffers?
A. They are made up of weak acid and its salts
B. They are made up of weak base and its salts
C. they tend to resist a change in pH value
D.) ihey are made up of strong acid and weak base

Q-21. Relative order of acidity of $\mathrm{HF}, \mathrm{HCl}, \mathrm{HBr}$ and HI acids is
A. $\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}>\mathrm{HF}$
B. $\mathrm{HF}>\mathrm{HCl}>\mathrm{HBr}>\mathrm{H}$
C) $\mathrm{HI}>\mathrm{HBr}>\mathrm{HCl}>\mathrm{HF}$
D. $\mathrm{HF}>\mathrm{HI}>\mathrm{HCl}>\mathrm{HBr}$

Q-22. The pH of 0.1 M HCl solution is
A. 0.01
B. 0.1
C. 0.2
D) 1


Q-23. Which of the following amino acid have two carboxylis groups?
A) Aspartic acid
B. Histidine
C. Gluatmine
D. none

Q-24. How many isomeric aldoses are possible for the molecular formula $\mathrm{C}_{6} \mathrm{H}_{12} \mathrm{O}_{6}$
A. 2
B. 4
C. 6
D. 16

Q-25. Epimers are compounds that differ in
A. functional group
B. configuration at alpha carbon
C. configuration at any carbon
D. none

Q-26. Select a basic amino acid
A. Glycine
B. Cystine
C. Alanine
D. Lysine

Q-27. Apoenzyme is
A. Hydrolytic enzyme
B. Oxidative enzyme
C. Coenzyme
D. Protein part of enzyme after removal of coenzyme

Q-28. A sugar present in DNA is
A. D-ribose

B. D-glucose
C. 2-Deoxy-D-Ribose
D. None

Q-29. Which vitamin is responsible for vision?
A. Retinol

B. Vit D
C. Tocopherols
D. Vit. K

Q-30. Which dt the following is periodic propery?
A. Atomic volume
B. Metallic character
C. Ionization energy
(D) All
B. n-pentadecane
D. n-hexadecane
A. n-hexane

Q-46. Cetane is
Q-31. Which of the following has largest size ?
A. Na*
B. Cl
C. F
D. $\mathrm{Fe}^{2+}$

Q-32. Which group contains elements that exist as monoatomic molecules?
A. 1
B. 2
C. 14
(D) 18
?

Q-33. Which of the following element is more metallic?
A. P
B. As
C. Sb
(D) Bi

Q-34. Which of the following factor effect ionization energy?
A. Atomic radius
B. Nuclear charge
C Shedding effect
(D) All
Q.35. Which of the following element is metalloid?
A. Pb
B. carbon
(C) As
D. Mg


Q-47. Which among the following is not nucleophile?
A. CH 3 NH 3
B. $\mathrm{CH}_{2} \mathrm{CH} 2$
C. $\mathrm{CH}_{3}+$

Q-48. Which of the following is strong acid?
A. Benzoic acid
B. m-nitrobenzoic acid
C. o-nitrobenzoic acid
D. p-nitrobenzoic acid

Q-49. Which of the following orbital does not make sense?
A. $6 f$
B. $4 f$
C. 7 s
D 5 g

Q-50. The maxium value of order of reaction

## may be?

A. 1
B. 2
C. 3
D. 4

Q-36. Which of the following compound does not following octet rule?
A. $\mathrm{CS}_{2}$
B. $\mathrm{PBr}_{3}$
C. IBr
(D) $\mathrm{BrF}_{5}$


Q-37. Which of the following halides have the lowest melting point?
A. NaCl
B. NaF
C. NaBr
(D) NaI

Q-38. Which of the following halides has zero
dipole moment?
A. $\mathrm{NH}_{3}$
B. $\mathrm{Cl}_{2}$
C. $\mathrm{NF}_{3}$
D. $\mathrm{CHCl}_{3}$


Q-39. The state of hybridization of carbon in COL S ?
A. sp
B. sp 2
C. sp 3
D. dsp 2

Q-40. Which one of the following is non-polar?
A CH2Cl
(B) $\mathrm{CCl}_{4}$
C. $\mathrm{CHCl}_{3}$
D. $\mathrm{CHCl}_{3}$

Q-41. Silvite is an ore of..?
A. Ca
B. Mg
C. Ba
D. K
Q-42. Which elements are non metals?
A $N \& P \quad$ B. As \& $S b$
C. $\mathrm{Sb} \& \mathrm{Bi}$
D. $\mathrm{Ba} \& \mathrm{Bi}$


Q-43. Among oxides of nitrogen all are gases
except
A. N2O5
B. N 2 O
C. NO
(D) $\mathrm{N}_{2} \mathrm{O}_{3}$

Q-44. All halogens form oxyacid except
A. Ftourine B. Chlorine C. Bromine D. Iodine
$X$

Q-45. Compunds HCN and HNC are
A. Tautomer
B. Metamers
(C) functional isomers
D. Conformers


## Screening test for the appointment of TGT (Contract) <br> Subject: Chemistry

Note: Choose the correct answer and all questions are compulsory.

Q-1. The process of identifying the component present in a sample is called
A) Quantitative analysis
B. Qualitative analysis
C. Volumetric analysis
D. Gravimetric analysis

Q-2. Conductometry is based on
A. Electric current
B. Electrical potential
C. Dielectric constant
D. Electrical conductance

Q-3. Which of the following technique is based on the emission of light radiation?

A. Flame photometry
B. Atomic absorption spectrophotometry
C. Raman spectroscopy
D. Conductometry

Q-4. Which of the following method is based on the solubility difference between the analyte and the unwanted components
A. Distillation
B. Complex formation
C. Electrodeposition
D. Precipitation

Q-5. Which of the following technique is based on the rate of reaction?
$A$
$B$
$C$
$C$
$C$
A. Isotopic dilution analysis
B. Flow injection analysis
C. Mass spectrometry
D. None

Q-6. The term accuracy refers to how near the observed value is to

## A. Mean value

B. Significant value
(C) Actual value
D. None

## Ullah Khan <br>  <br>  <br> 

Q-7. When HCl is titrated against NaOH the pH at equivalence point will be
A. None
B. equal to 7
C. less than 7
D. greater than 7


Q-8. When CH 3 COOH is titrated against NaOH the pH at the end point is
A. None
B. equal to 7
C. less than 7
D. greater than 7

Q-9. Which of the following is best indicator for titration of $\mathrm{NH}_{4} \mathrm{OH}$ with HCl ?
A. Methyl red
B) Methyl orange
C. Phenolphthalein
D. None

Q-10. Soap is a salt of compound known as
A. Acetic acid
B. Formic acid
C) Fatty acid
D. Amino acid

Q-11. The colour of water in a take is due to
A. Excessive growth of sea weeds
B) Algae
C. Grass
D. Other pollution

Q-12. Cement sa mixture of
A. Clay and clinker
B. Clay, limestone and gypsum
C. Limestone and gypsum
D. Limestone and clay

Q-13. The percentage of nitrogen in urea is
A. $36 \%$
B. $46 \%$
C. $56 \%$
D. $66 \%$

Q-14. The percentage of nitrogen in ammonia is
A. $32 \%$
B. $42 \%$
C. $72 \%$
D. $82 \%$



Q-15. Copper occur in nature as
A. Native
B. Combined
C. Both
D. None

Q-16. The formula of Cryolite is
A. $\mathrm{Na}_{3} \mathrm{AlF}_{3}$
B. $\mathrm{Na}_{3} \mathrm{AlF}_{4}$
C. $\mathrm{Na}_{3} \mathrm{AlF}_{5}$
(D) None

Q-17. Cells with nuclear envelop belongs to
A. Prokaryotes
B. Eukaryotes
C. Both
D. None


Q-18. The diameter of a typical plant or animal cell is
A. 1 to $5 \mu \mathrm{~m}$
B) 5 to $10 \mu \mathrm{~m}$
C. 10 to $50 \mu \mathrm{~m}$
D. 5 to $100 \mu \mathrm{~m}$


Q-19. Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by
A) Davy
B. Liebig
C. Boyle
D. Rouelle


Q-20. Which of the following statement is not true regarding buffers?
A. They are made up of weak acid and its salts B. They are made up of weak base and its salts
C. they tend to resist a change in pH value D. they are made $\mu$ of strong acid and weak base

Q-21. Relative order of acidity of $\mathrm{HF}, \mathrm{HCl}, \mathrm{HBr}$ and HI acids is
A. $\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}>\mathrm{HF}$
B. $\mathrm{HF}>\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}$
C) $\mathrm{H}>\mathrm{HBr}>\mathrm{HCl}>\mathrm{HF}$
D. $\mathrm{HF}>\mathrm{HI}>\mathrm{HCl}>\mathrm{HBr}$

Q-22. The pH of 0.1 M HCl solution is
A. 0.01
B. 0.1
C. 0.2
D. 1

Q-23. Which of the following amino acid have two carboxylic groups?
A) Aspartic acid
B. Histidine
C. Gluatmine
D. none

Q-24. How many isomeric aldoses are possible for the molecular formula $\mathrm{C}_{6} \mathrm{H}_{12} \mathrm{O}_{6}$
(A) 2
B. 4
C. 6
D. 16


Q-25. Epimers are compounds that differ in
A. functional group
B. configuration at alpha carbon
C. configuration at any carbon
D. none

Q-26. Select a basic amino acid
A. Glycine

B. Cystine
C. Alanine
D. Lysine

Q-27. Apoenzyme is
A. Hydrolytic enzyme
B. Oxidative enzyme
C. Coenzyme
(D) Protein part of enzyme after removal of coenzyme

## Q-28. A sugar present in DNA is

A. D-ribose
B. D-glucose
C. 2-Deoxy-D-Ribose
D. None

Q-29. Which vitamin is responsible for vision?
(A) Retinol
B. Vil D
C. Tocopherols
D. Vit. K

Q-30. Which of the following is periodic propery?
A. Atomic volume
B. Metallic character
C. Ionization energy
D. All

Q-31. Which of the following has largest size?
A. $\mathrm{Na}^{+}$
B. $\mathrm{Cl}^{-}$
C. $F$
(D.) $\mathrm{Fe}^{2+}$

Q-32. Which group contains elements that exist
as monoatomic molecules?
A. 1
B. 2
C. 14

D. 18

Q-46. Cetane is

A. n-hexane
B. n-pentadecane
C. n-octane
D. n-hexadecane

Q-47. Which among the following is not nucleophile?
A. $\mathrm{CH}_{3} \mathrm{NH}_{3}$
B. $\mathrm{CH}_{2} \mathrm{CH}_{2}$
(C)

Q-33. Which of the following element is more
metallic?
A. P
B. As

D. Bi

Q-34. Which of the following factor effect ionization energy?

A. Atomic radius
B. Nuclear charge
C. Shedding effect
(D.) All

Q-48. Which of the following is strong acid?

## A. Benzoic acid

wing is strong acid?
B. m-nitrobenzoic acid
D. p-nitrobenzoic acid
C. O-nitrobenzoic acid

Q-49. Which of the following orbital does not make sense?
A. $6 f$
B. $4 f$
C. 7 s
D) 5 g

Q-50. The maximum value of order of reaction metalloid?
A. Pb
B. carbon
(C) As
D. Mg

## may be?

A. 1
B. 2
( 5
D. 4

Q-36. Which of the following compound does not following octet rule?
A. $\mathrm{CS}_{2}$
B. $\mathrm{PBr}_{3}$
C. IBr
(D.) $\mathrm{BrF}_{5}$

Q-37. Which of the following halides have the lowest melting point?
A. NaCl
B. NaF
C. NaBr
DNa l

Q-38. Which of the following halides has zero dipole moment?
A. $\mathrm{NH}_{3}$
B. $\mathrm{Cl}_{2}$
C. $\mathrm{NF}_{3}$
D. $\mathrm{CHCl}_{3}$ Х

Q-39. The state of hybridization of carbon in
CO 2 is?
A. sp
B. sp 2
C. sp 3
D. dsp2

$0=0$
$2^{2}$

Q-40. Which one of the following is non-polar?
A. $\mathrm{CH}_{2} \mathrm{Cl}$
(B) $\mathrm{CCl}_{4}$
C. $\mathrm{CHCl}_{3}$
D. $\mathrm{CHCl}_{3}$

Q-41. Silvite is an ore of..?
(A) Ca
B. Mg
C. Ba
D.K $X$

Q-42. Which elements are non metals?
A. $N \& P$
B. As \& Sb
C. $\mathrm{Sb} \& \mathrm{Bi}$
D. $\mathrm{Ba} \& \mathrm{Bi}$
Q-43. Among oxides of nitrogen all are gases except
A. N 2 O 5
B. N 2 O
C. NO NOB


Q-44. All halogens form oxyacid except
A. Flourine B. Chlorine
C. Bromine D D Iodine $X$

Q-45. Compunds HCN and HNC are
A. Tautomers
B. Metamer
C. functional isomers
D. Conformers


## Screening test for the appointment of TGT (Contract)

Subject: Chemistry
Note: Choose the correct answer and all questions are compulsory

Q-1. The process of identifying the component present in a sample is called

Quantitative analysis
B. Qualitative analysis
C. Volumetric analysis
-D. Gravimetric analysis

## Q-2. Conductometry is based on

A. Electric current
B. Electrical potential
C. Dielectric constant
D. Electrical conductance

Q-3. Which of the following technique is based on the emission of light radiation?

A. Flame photometry

- Atomic absorption spectrophotometry
C. Raman spectroscopy
- D. Conductometry

Q-4. Which of the following method is based on the solubility difference between the analyse and the unwanted components
A. Distillation
B. Complex formation
C. Electrodeposition
(D) Precipitation

Q-5. Which of the following technique is based on the rate of reaction?
A. Isotopic dilution analysis
B) Flow injection analysis
C. Mass spectrometry
D. None

Q-6. The term accuracy refers to how near the observed value is to
A. Mean value
B. Significant value
C. Actual value
D. None

Q-7. When HCl is titrated against NaOH the pH
at equivalence point will be
A None
B. equal to 7

- C. less than 7
D. greater than 7

Q-8. When CH 3 COOH is titrated against NaOH the pH at the end point is
A. None
B. equal to 7
C. less than 7
D) greater than 7

Q-9. Which of the following is best indicator for titration of $\mathrm{NH}_{4} \mathrm{OH}$ with HCl ?

A. Methyl red
B. Methyl orange
C) Phenolphthaleine
D. None

Q-10. Soap is a salt of compound known as
A. Acetic acid
B. Formic acid
C. Fatty acid
D. Amino acid

Q-11. The colour of water in a lake is due to
A Excessive growth of sea weeds
B. Algae
C. Grass
D. Other pollution

Q-12. Cement sa mixture of
A. Clay and clinker
B. Clay, limestone and gypsum
C. Limestone and gypsum
D. Limestone and clay

Q-13. The percentage of nitrogen in urea is
A. $36 \%$
$\times \begin{aligned} & \text { B } 46 \% \\ & 56 \%\end{aligned}$
D. $66 \%$

Q-14. The percentage of nitrogen in ammonia is
A. $32 \%$


110 age


Q-16. The formula of Cryolite is


Q-17. Cells with nuclear envelop belongs to


Q-18. The diameter of a typical plant or animal cell is
A. 1 to $5 \mu \mathrm{~m}$
B. 5 to $10 \mu \mathrm{~m}$
C. 10 to $50 \mu \mathrm{~m}$
(D. 5 to $100 \mu \mathrm{~m}$

Q-19. Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by

A. Davy
B. Liebig
C. Boyle
D. Rouelle

Q-20. Which of the following statement is not true regarding buffers?
A. They are made up of weak acid and its salts B. They are made up of weak base and its salts C) they tend to resist a change in pH value D. they are made up of strong acid and weak base

Q-21. Relative order of acidity of $\mathrm{HF}, \mathrm{HCl}, \mathrm{HBr}$ and Hl acids is
A. $\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}>\mathrm{HF}$
B. $\mathrm{HF}>\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}$
C. $\mathrm{HI}>\mathrm{HBr}>\mathrm{HCl}>\mathrm{HF}$
D. $\mathrm{HF}>\mathrm{HI}>\mathrm{HCl}>\mathrm{HBr}$
$>A^{Q-22.01}$
Q-22. The pH of 0.1 M HCl solution is
B. 0.1
C. 0.2
D. 1

Q-23. Which of the following amino acid have two carboxylic groups?
A. Aspartic acid


Q-24. How many isomeric aldoses are possible for the molecular formula $\mathrm{C}_{6} \mathrm{H}_{12} \mathrm{O}_{6}$
A. 2


Q-25. Epimers are compounds that differ in
(A) functional group
B. configuration at alpha carbon
C. configuration at any carbon
D. none

Q-26. Select a basic amino acid
A. Glycine
C. Alanine
D. Lysine

Q-27. Apoenzyme is
A. Hydrolytic enzyme
B. Oxidative enzyme
C. Coenzyme
D. Protein part of enzyme after removal of coenzyme

Q-28. A sugar present in DNA is
A D-ribose
B) D-glucose
C. 2-Deoxy-D-Ribose
D. None

Q-29. Which vitamin is responsible for vision?
A. Retinol
B. Vit D

G. Tocopherots

Q-30. Which of the following is periodic propery?
$\begin{array}{ll}\text { A. Atomic volume } & \text { B. Metallic character } \\ \text { C. Ionization energy } & \text { D. All }\end{array}$

$$
\underbrace{n^{3}}
$$

Q-31. Which of the following has largest size? A. $\mathrm{Na}^{+} \quad$ B. Cl C.F $\quad$ (D.) $\mathrm{Fe}^{2+}$

Q-32 Which group contains elements that exist as monoatomic molecules?
A. 1
B. 2
C. 14


Q-33. Which of the following element is more metallic?
A. P
B. As
C. Sb
(D) Bi

Q-34. Which of the following factor effect ionization energy?
A. Atomic radius
B. Nuclear charge
C. Shedding effect
(D) All

Q-35. Which of the following element is metalloid?

- A. Pb
B. carbon
C. As
D. Mg

Q-36. Which of the following compound does not following octet rule?
A. $\mathrm{CS}_{2}$
B. $\mathrm{PBr}_{3}$
(C) Br
D. $\mathrm{BrF}_{5}$

Q-37. Which of the following halides have the
lowest melting point?
A. NaCl
B. NaF
C. NaBr
D. Nal

Q-38. Which of the following halides has zero dipole moment?
A. $\mathrm{NH}_{3}$
(B) $\mathrm{Cl}_{2}$
C. $\mathrm{NF}_{3}$
D. $\mathrm{CHCl}_{3}$

Q-39. The state of hybridization of carbon in CO 2 is?
(A) sp
B. sp 2
C. sp 3
D. dsp 2

Q-40. Which one of the following is non-polar?
A. $\mathrm{CH}_{2} \mathrm{Cl}$
B) $\mathrm{CCl}_{4}$
C. $\mathrm{CHCl}_{3}$
D. $\mathrm{CHCl}_{3}$

Q-41. Silvite is an ore of..?
$\begin{array}{lll}\text { A. Ca B. } \mathrm{Mg} & \text { C. Ba D.K }\end{array}$
Q-42. Which elements are non metals?
$X$
A. N \& P
B. As \& Sb
C. Sb \& $\mathrm{Bi} \quad \mathrm{D} . \mathrm{Ba} \& \mathrm{Bi}$
Q-43. Among oxides of nitrogen all are gases
except
A. N 2 O 5 B 2 O
C. NO
D. N 2 O 3

Q-44. All halogens form oxyacid except
A. Flourine B. Chlorine C. Bromine D. Iodine

Q-45. Compounds HCN and HNC are

[^0]c.


Name: Rozina
F/ name: Sher Daraz Khan

## Screening test for the appointment of TGT (Contract)

 Subject: ChemistryNote: Choose the correct answer and all questions are compulsory.

Q-1. The process of identifying the component present ind sample is called
A. Quantitative analysis
(B.) Qualitative analysis
C. Volumetric analysis
D. Gravimetric analysis

Q-2. Conductometry is based on
A. Electric current
B. Electrical potential
E. Dielectric constant
(D.) Electrical conductance

Q-3. Which of the following technique is based on the emission of light radiation?
A. Flame photometry
B. Atomic absorption spectrophotometry
C. Ramen spectroscopy
D. Conductometry

Q-4. Which of the following method is based on the solubility difference between the analyte and the unwanted components
A. Distillation
B. Complex formation

厄. Electrodeposition
D. Precipitation

Q-5. Which of the following technique is based on the rate of reaction?
A. Isotopic dilution analysis
B. Flow injection analysis
C. Mass spectrometry
D. None

Q-6. The term accuracy refers to how near the observed value is to
A. Mean value
B. Significant value
(C) Actual value
D. None

Q-7. When H (1 is titrated against NaOH the pH
at equivalence point will be

## A. None

## (B) equal to 7

C. less than 7
D. greater than 7

Q-8. When CH 3 COOH is titrated against NaOH the pH at the end point is
VA. None

B. equal to 7
C. less than 7
D. greater than 7

Q-9. Which of the following is best indicator for titration of $\mathrm{NH}_{4} \mathrm{OH}$ with HCl ?
A) Methyl red
B. Methyl orange
C. Phenolphthalein
D. None

Q-10. Soap is a salt of compound known as
A. Acetic acid
8. Formic acid
C. Fatty acid
D. Amino acid

Q-11. The colour of water in a lake is due to
(A) Excessive growth of sea weeds
B. Algae
C. Grass
D. Other pollution

Q-12. Cement sa mixture of
A. Clay and clinker
B. Clay, limestone and gypsum
-C. Limestone and gypsum
D. Limestone and clay

Q-13. The percentage of nitrogen in urea is
A. $36 \%$
C. $56 \%$
D. $66 \%$
Q- 14. The percentage of nitrogen in ammonia is
A. $32 \%$
C. $42 \%$
D. $82 \%$

5
D. $66 \%$

Q-15. Copper occur in nature as
A. Native
B. Combined
C. Both
D. None


Q-16. The formula of Cryolite is
A. $\mathrm{Na}_{3} \mathrm{AlF}_{3}$
B. $\mathrm{Na}_{3} \mathrm{AlF}_{4}$
C. $\mathrm{Na}_{3} \mathrm{AlF}_{5}$
(D) None


Q-17. Cells with nuclear envelop belongs to
A. Prokaryotes
B) Eukaryotes
C. Both
D. None

Q-18. The diameter of a typical plant or animal cell is
(A.) 1 to $5 \mu \mathrm{~m}$
B. 5 to $10 \mu \mathrm{~m}$
C. 10 to $50 \mu \mathrm{~m}$
D. 5 to $100 \mu \mathrm{~m}$


Q-19. Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by
(A.) Davy
B. Liebig

C. Boyle
D. Rouelle

Q-20. Which of the following statement is not true regarding buffers?
A. They are made up of weak acid and its salts
B. They are made up of weak base and its salts
C. they tend to resist a change in pH value
D. they are made upef strong acid and weak base

Q-21. Relative order of acidity of $\mathrm{HF}, \mathrm{HC}, \mathrm{HBr}$ and HI acids is
A. $\mathrm{HCl}>\mathrm{HBr}>\mathrm{Hl}>\mathrm{HF}$
B. $\mathrm{HF}>\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}$
C. $\mathrm{HI}>\mathrm{HBr}>\mathrm{HCl}>\mathrm{HF}$
D. $\mathrm{HF}>\mathrm{HI}>\mathrm{HCl}>\mathrm{HBr}$

Q-22. The pH of 0.1 M HCl solution is

A. 0.01
B. 0.1
C. 0.2
D. 1

Q-23. Which of the following amino acid have two carboxylic groups?
A) Aspartic acid

B. Histidine
C. Gluatmine
D. none

Q-24. How many isomeric aldoses are possible for the molecular formula $\mathrm{C}_{6} \mathrm{H}_{12} \mathrm{O}_{6}$
(A.) 2
B. 4
C. 6
D. 16


Q-25. Epimers are compounds that differ in
A functional group
(B.) configuration at alpha carbon
C. configuration at any carbon
D. none

Q-26. Select a basic amino acid
A. Glycine
B. Cystine
C. Alanine
D. Lysine


Q-27. Apoenzyme is
A. Hydrolytic enzyme
B. Oxidative enzyme
C. Coenzyme
(D.) Protein part of enzyme after removal of coenzyme

Q-28. A sugar present in DNA is


Q-29. Which vitamin is responsible for vision?
A. Retinol
B. Vit D
C. Tocopherols
D. Vit. K

Q-30. Which of the following is periodic propery?
A. Atomic volume Metallic character
C. Ionization energy All
C. Ionization energy (D) All

Q-31. Which of the following has largest size ?
$\begin{array}{llll}\text { A. } \mathrm{Na}^{+} & \text {(B. } \mathrm{Cl}^{-} & \text {C. } \mathrm{F}^{-} & \text {D. } \mathrm{Fe}^{2+}\end{array}$
Q-32. Which group contains elements that exist as monoatomic molecules?
B. 2
C. 14
(D.) 18

## A. 1

exist


Q-46. Cetane is
A. n-hexane
C. n-octane
(B.) n-pentadecane
D. n-hexadecane

Q-47. Which among the following is not nucleophile?
A CH 3 NH 3
(C.) $\mathrm{CH}_{3}+$
B. $\mathrm{CH}_{2} \mathrm{CH}_{2}$
D. OH


Q-33. Which of the following element is more metallic?
A. P
B. As
C. Sb
(D.) Bi

Q-34. Which of the following factor effect ionization energy?
A. Atomic radius
C. Shiedding effect
B. Nuclear charge

Q-35. Which of the following element is metalloid?
A. Pb
B. carbon
(C) As
D. Mg
Q-48. Which of the following is strong acid?
A. Benzoic acid
(B.) m-nitrobenzoic acid
C. o-nitrobenzoic acid
D. p-nitrobenzoic acid

Q-49. Which of the following orbital does not make sense?
A. $6 f$
B. $4 f$
C. 7 s
(D) $5 g$

Q-50. The maximum value of order of reaction may be?
A. 1
B. 2
(C.) 3
D. 4

Q-36. Which of the following compound does not following octet rule?
A. $\mathrm{CS}_{2}$
B. $\mathrm{PBr}_{3}$
C. IBr
(D) $\mathrm{BrF}_{5}$

5
Q-37. Which of the following halides have the lowest melting point?
A. NaCl
B. NaF
C. NaBr
(D) NaI

Q-38. Which of the following halides has zero dipole moment?
A. $\mathrm{NH}_{3}$
B. $\mathrm{Cl}_{2}$
(C) $\mathrm{NF}_{3}$
D. $\mathrm{CHCl}_{3}$


Q-39. The state of hybridization of carbon in
CO 2 is?
A. sp
(B) sp 2
C. sp 3
D. dsp 2


Q-40. Which one of the following is non-polar?
A. $\mathrm{CH}_{2} \mathrm{Cl}$
(B.) $\mathrm{CCl}_{4}$
C. $\mathrm{CHCl}_{3}$
D. $\mathrm{CHCl}_{3}$


Q-41. Silvite is an ore of..?
A. Ca
B. Mg
C. Ba
(D.) K
$>$
Q-42. Which elements are non metals?
(A) $N \& P$
B. As \& Sb
C. Sb \& Bi
D. Ba \& Bi
$X$

Q-43. Among oxides of nitrogen all are gases
except
A. NO
(B) N 2 O
C. NO
D. N 2 O 3
$X$

Q-44. All halogens form oxyacid except
A. Flouring B. Chlorine C. Bromine (D.) Iodine


Q-45. Compunds HCN and HNC are
A. Tautomer
B. Metamer
C. functional isomers
D. Conformers


Name = Caiseo Khan
F/Name Ayer Shan

## Screening test for the appointment of TGT (Contract)

 Subject: ChemistryNote: Choose the correct answer and all questions are compulsory.

Q-1. The process of identifying the component present in a sample is called
A. Quantitative analysis
(B) Qualitative analysis
C. Volumetric analysis
D. Gravimetric analysis

Q-2. Conductometry is based on
A. Electric current
B. Electrical potential
C. Dielectric constant
(D) Electrical conductance

Q-3. Which of the following technique is based on the emission of light radiation?

A Flame photometry
B. Atomic absorption spectrophotometry
C. Raman spectroscopy
D. Conductometry

Q-4. Which of the following method is based on the solubility difference between the analyte and the unwanted components
A. Distillation
B. Complex formation
C. Electrodeposition
(D) Precipitation

Q-5. Which of the following technique is based on the rate of reaction?
A. Isotopic dilution analysis
B.) Flow injection analysis
C. Mass spectrometry
D. None

Q-6. The term accuracy refers to how near the observed value is to
A. Mean value
B. Significant value
C. Actual value
D. None

Q-7. When HC (s) titrated against $\sqrt{ } \mathrm{OOH}$ the pH at equivalence point will be
A. Hone
B) equal to 7
C. less than 7
D. greater than 7

Q-8. When CH 3 COOH is titrated against NaOH the pH at the end point is
A. None

B. equal to 7
C. less than 7
(D) greater than 7

Q-9. Which of the following is best indicator for titration of $\mathrm{NH}_{4} \mathrm{OH}$ with HCl ?


Q-10. Soapis a salt of compound known as
A. Acetic acid
B. Formic acid
C. Fatty acid
D. Amino acid

Q-11. The colour of water in a lake is due to
A) Excessive growth of sea weeds
B. Algae
C. Grass
D. Other pollution

Q-12. Cement s a mixture of
A. Clay and clinker
B. Clay, limestone and gypsum
C. Limestone and gypsum
D. Limestone and clay

Q-13. The percentage of nitrogen in urea is
A. $36 \%$
B. $46 \%$
. $56 \%$
D. $66 \%$

Q-14. The percentage of nitrogen in ammonia is
A. $32 \%$
B. $42 \%$
C. $72 \%$
D. $82 \%$


Q-15. Copper occur in nature as
A. Native
B. Combined
C. Both
D. None


Q-16. The formula of Cryolite is
A. $\mathrm{Na}_{3} \mathrm{AlF}_{3}$
(B) $\mathrm{Na}_{3} \mathrm{AlF}_{4}$
C. $\mathrm{Na}_{3} \mathrm{AlF}_{5}$
D. None

Q-17. Cells with nuclear envelop belongs to
A. Prokaryotes
B) Eukaryotes
C. Both
D. None

Q-18. The diameter of a typical plant or animal cell is
A. 1 to $5 \mu \mathrm{~m}$
B. 5 to $10 \mu \mathrm{~m}$
C. 10 to $50 \mu \mathrm{~m}$
D. 5 to $100 \mu \mathrm{~m}$


Q-19. Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by
A. Davy
(B. Liebig
C. Boyle

D. Rouelle

Q-20. Which of the following statement is not true regarding buffers?
A. They are made up of weak acio and its salts
B. They are made up of weak base and its salts
C. they tend to resist a change in pH value
(D) they are made upef strong acid and weak base

Q-21. Relative order of acidity of $\mathrm{HF}, \mathrm{HCH}, \mathrm{HBr}$ and Hl acids is
A. $\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}>\mathrm{HF}$
B. $\mathrm{HF}>\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}$
C. $\mathrm{HI}>\mathrm{HBr}>\mathrm{HCl}>\mathrm{HF}$
D. $\mathrm{HF}>\mathrm{HI}>\mathrm{HCl}>\mathrm{HBr}$

Q-22. The pH of 0.1 M HCl solution is
A. 0.01
B. 0.1
C. 0.2
D. 1

Q-23. Which of the following amino acid have two carboxylic groups?
A. Aspartic acid
B. Histidine
C. Gluatmine
D. none


Q-24. How many isomeric aldoses are possible for the molecular formula $\mathrm{C}_{6} \mathrm{H}_{12} \mathrm{O}_{6}$
A. 2
$\begin{aligned} & \text { B. } 4 \\ & \text { (C). } 6 \\ & \text { D. } 16\end{aligned}>$
Q-25. Epimers are compounds that differ in
A. functional group
(B.) configuration at alpha carbon
C. configuration at any carbon
D. none

Q-26. Select a basic amino acid
A. Glycine
B. Cystine
C. Alanine
D. Eysine


Q-27. Apoenzyme is
A. Hydrolytic enzyme
B. Oxidative enzyme
C) Coenzyme

D. Protein part of enzyme after removal of coenzyme

Q-28. A sugar present in DNA is
A. D-ribose
B. D-glucose
C. 2-Deoxy-D-Ribose
D. None

Q-29. Which vitamin is responsible for vision?
(A) Retinol

B. Vit D
C. Tocopherols
D. Vit. K

Q-30. Which of the following is periodtc propery?
A. Atomic volume B. Metallic character
C. Ionization energy
(D) All

Q-31. Which of the following has largest size?
A. $\mathrm{Na}^{+}$
B. $\mathrm{Cl}^{-}$
C. $\mathrm{F}^{-}$
D. $\mathrm{Fe}^{2+}$


Q-46. Cetane is

A. n-hexane
B. n-pentadecane
C. n-octane
D. n-hexadecane

Q-32. Which group contains elements that exist as monoatomic molecules?
A. 1
B. 2
C. 14
(D) 18

Q-47. Which among the following is not nucleophile?
A. CH 3 NH 3
B. $\mathrm{CH}_{2} \mathrm{CH} 2$ CH+
D. OH

Q-33. Which of the following element is more metallic?
A. P
B. As
C. Sb
(D) Bi

Q-34. Which of the following factor effect ionization energy?
A. Atomic radius
C. Shiedding effect
B. Nuclear charge

Q-35. Which of the following element is metalloid?
A. Pb
B. carbon
C. $A s$
D. Mg


Q-48. Which of the following is strong acid?
A. Benzoic acid
C. o-nitrobenzoic acid
B. m-nitrobenzoic acid

Q-49. Which of the following orbital does not make sense?
A. $6 f$
B. 4 f
C. 7 s
(D) 5 g

Q-50. The maxium value of order of reaction may be?
A. 1
B. 2
(C) 3
D. 4

Q-36. Which of the following compound does not following octet rule?
A. $\mathrm{CS}_{2}$
B. $\mathrm{PBr}_{3}$
C. IBr
(D) $\mathrm{BrF}_{5}$

Q-37. Which of the following halides have the lowest melting point?
A. NaCl
B. NaF
C. NaBr
(D) Nail

Q-38. Which of the following halides has zero dipole moment?
A. $\mathrm{NH}_{3}$
B. $\mathrm{Cl}_{2}$
C. $\mathrm{NF}_{3}$
D. $\mathrm{CHCl}_{3}$

Q-39. The state of hybridization of carbon in
 CO 2 is?
(A) sp
B. sp 2
C. $s p 3$
D. dsp 2

$$
\begin{array}{ll}
8=8 & 8=5 \\
4=5 & 8=7
\end{array}
$$

Q-40. Which one of the following is non-polar?

$$
8 x^{2}+8
$$

A. $\mathrm{CH}_{2} \mathrm{Cl}$
B. $\mathrm{CCl}_{4}$
C. $\mathrm{CHCl}_{3}$
D. $\mathrm{CHCl}_{3}$

Q-41. Silvite is an ore of..?
A. Ca
B. Mg
C. Ba
D. K

Q-42. Which elements are non metals?
(A) $N \& P$
B. As \& Sb
C. $\mathrm{Sb} \& \mathrm{Bi} \quad \mathrm{D} . \mathrm{Ba} \& \mathrm{Bi}$


Q-43. Among oxides of nitrogen all are gases
except
A. N 2 O 5
B. N 2 O
C. NO
D. N 2 O 3

$$
y
$$

Q-44. All halogens form oxyacid except

$$
M>s
$$

Q-45. Compounds HCN and HNC are
A. Tautomer
C. functional isomers
D. Conformers


## A. Flouring B. Chlorine C) Bromine D. Iodine $X$



$$
\begin{array}{ll}
3 \\
8 & * \\
8 & *
\end{array}
$$



$$
1<=5
$$



$$
\begin{aligned}
& \text { MAHEEN RAHIM } \\
& 17301.9730205-4
\end{aligned}
$$

## Screening test for the appointment of TGT (Contract)

## Subject: Chemistry

Note: Choose the correct answer and all questions are compulsory

Q-1. The process of identifying the component present in a sample is called
(A) Quantitative analysis
B. Qualitative analysis
C. Volumetric analysis
D. Gravimetric analysis

Q-2. Conductometry is based on
A. Electric current
B. Electrical potential
C. Dielectric constant
D. Electrical conductance

Q-3. Which of the following technique is based on the emission of light radiation?
A. Flame photometry
B. Atomic absorption aspect photometry
C. Raman spectroscopy
D. Conductometry

Q-4. Which of the following method is based on the solubility difference between the analyte and the unwanted components
A. Distillation
B. Complex formation
C. Electrodeposition
(D) Precipitation

Q-5. Which of the following technique is based on the rate of reaction?
A. Isotopic dilution analysis
B. Flow injection analysis
C. Mass spectrometry
(D. None

Q-6. The term accuracy refers to how near the observed value is to


Q-7. When HCl is titrated against NaOH the pH at equivalence point wit be
A. None
B. equal to 7
C. less than 7
D. greater than 7


Q-8. When CH 3 COOH is titrated against NaOH the pH at the end point is

A. None
B. equal to 7
C. Jess than 7
D. greater than 7

Q-9. Which of the following is best indicator for titration of $\mathrm{NH}_{4} \mathrm{OH}$ with HCl ?
A. Methyl red
B. Methyl orange
C. Phenolphthalein
D. None

Q-10. Soap is a salt of compound known as
A. Acetic acid
B. Formic acid
C. Fatty acid
D. Amino acid

Q-11. The colour of water in a lake is due to
(A) Excessive growth of sea weeds
B. Algae
C. Grass
D. Other pollution

Q-12. Cement sa mixture of
A. Clay and clinker
B. Clay, limestone and gypsum
C. Limestone and gypsum
D. Limestone and clay

Q-13. The percentage of nitrogen in urea is
A. $36 \%$
B. $46 \%$
C. $56 \%$
D. $66 \%$
Q-14. The percentage of nitrogen in ammonia is
A. $32 \%$
B. $42 \%$
C. $72 \%$
(D. $82 \%$


Q-15. Copper occur in nature as
A. Native

B Combined
(C. Both
D. None

Q-16. The formula of Cryolite is
A. $\mathrm{Na}_{3} \mathrm{AlF}_{3}$
B. $\mathrm{Na}_{3} \mathrm{AlF} \mathrm{F}_{4}$
C. $\mathrm{Na}_{3} \mathrm{AlF}_{5}$
D. None

Q-17. Cells with nuclear envelop belongs to
A. Prokaryotes
B. Eukaryotes
C. Both
D. None

Q-18. The diameter of a typical plant or animal cell is
A. 1 to $5 \mu \mathrm{~m}$
B. 5 to $10 \mu \mathrm{~m}$
C. 10 to $50 \mu \mathrm{~m}$
(D) 5 to $100 \mu \mathrm{~m}$


Q-19. Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by
(A) Davy
B. Liebig
C. Boyle

D. Rouelle

Q-20. Which of the following statement is not true regarding buffers?
A. They are made up of weak acid and its salts
B. They are made up of weak base and its salts
C. they tend to resist a change in pH value
(D. they are made up of strong acid and weak base

Q-21. Relative order of acidity of $\mathrm{HF}, \mathrm{HCl}, \mathrm{HBr}$ and Hl acids is
A. $\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}>\mathrm{HF}$
B. $\mathrm{HF}>\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}$
C. $\mathrm{HI}>\mathrm{HBr}>\mathrm{HCl}>\mathrm{HF}$
D. $\mathrm{HF}>\mathrm{Hl}>\mathrm{HCl}>\mathrm{HBr}$

Q-22. The pH of 0.1 M HCl solution is
A. 0.01
B. 0.1
C. 0.2 (D. 1

Q-23. Which of the following amino acid have two carboxylic groups?


Q-24. How many isomeric aldoses are possible for the molecular formuta $\mathrm{C}_{6} \mathrm{H}_{12} \mathrm{O}_{6}$
(A) 2
B. 4
C. 6
D. 16

Q-25. Epimers are compounds that differ in
A. functional group
B. configuration at alpha carbon
(C. configuration at any carbon
D. none

Q-26. Select a basic amino acid
A. Glycine
B. Cystine
C. Alanine
(D. Lysine


Q-27. Apoenzyme is
A. Hydrolytic enzyme
B. Oxidative enzyme
C. Coenzyme
(D. Protein part of enzyme after removal of coenzyme

Q-28. A sugar present in DNA is
A. D-ribose
B. D-glucose
C. 2-Deoxy-D-Ribose
D. None

Q-29. Which vitamin is responsible for vision?



Q-30. Which of the following is periodic propery?
A. Atomic volume
B. Metatic character
C. Ionization energy

Q-31. Which of the following has largest size?
(A. $\mathrm{Na}^{+}$
B. Cr
C. $F^{-}$
D. $\mathrm{Fe}^{2+}>$

Q-46. Cetane is
(A )-hexane
B. n-pentadecane
C. n-octane
D. n-hexadecane

Q-32. Which group contains elements that exist as monoatomic molecules?
A. 1
B. 2
(C) 14
D. 18
$X$
Q-33. Which of the following element is more metallic?
A. P
B. As
C. Sb
(D) Bi

Q-34. Which of the following factor effect ionization energy?

Q-47. Which among the following is not nucleophile?
A. CH 3 NH 3
B. $\mathrm{CH}_{2} \mathrm{CH}_{2}$
C. $\mathrm{CH}_{3}+$
D. OH

Q-48. Which of the following is strong acid?
A. Benzoic acid
B. m-nitrobenzoic acid
C. o-nitrobenzoic acid
D. p-nitrobenzoic acid

Q-49. Which of the following orbital does not make sense?
A. $6 f$
B. $4 f$
C. 7 s
D 5 g

Q-50. The maxium value of order of reaction may be?
A. 1
B. 2
(C) 3
D. 4

A. Atomic radius
C. Shiedding effect
nuclear charge

Q-35. Which of the following element is
metalloid?
A. Pb
B. carbon
C. As
D. Mg


Q-36. Which of the following compound does not following octet rule?
A. $\mathrm{CS}_{2}$
B. $\mathrm{PBr}_{3}$
C. IBr
D. $\mathrm{BrF}_{5}$


Q-37. Which of the following halides have the
lowest melting point?
A. NaCl
B. NaF
C. NaBr
(D) NaI

Q-38. Which of the following halides has zero dipole moment?
A. $\mathrm{NH}_{3}$
(B. $\mathrm{Cl}_{2}$
C. $\mathrm{NF}_{3}$
D. $\mathrm{CHCl}_{3}$

Q-39. The state of hybridization of carbon in

CO 2 is?
A. $s p$
B. sp 2
sp 3
D. dsp 2


Q-40. Which one of the following is non-polar?
A. $\mathrm{CH}_{2} \mathrm{Cl}$
(B) $\mathrm{CCl}_{4}$
C. $\mathrm{CHCl}_{3}$
D. $\mathrm{CHCl}_{3}$

Q-41. Silvite is an ore of.,?
A. Ca
B. Mg
C. Ba
D. K

Q-42. Which elements are non metals?
$\begin{array}{llll}\text { A. N \& P } & \text { B. As \& Sb } & \text { C. } \mathrm{Sb} \text { \& } \mathrm{Bi} & \mathrm{D} . \mathrm{Ba} \& \mathrm{Bi}\end{array}$


Q-43. Among oxides of nitrogen all are gases except
A. N2O5
B. N 2 O
(C.) NO
D. N 2 O 3

Q-44. All halogens form oxyacid except
A. Flouring B. Chlorine C. Bromine D. Iodine


Q-45. Compunds HCN and HNC are
A. Tautomers
B. Metamer
(C. functional isomers
D. Conformers
agma Nomen

F/N Momqin Klan

## Screening test for the appointment of TGT (Contract)

Subject: Chemistry
Note: Choose the correct answer and all questions are compulsory.

Q-1. The process of identifying the component present in a sample is called
(A) Quantitative analysis
B. Qualitative analysis
C. Volumetric analysis
D. Gravimetric analysis

Q-2. Conductometry is based on
$X$
A. Electric current
B. Electrical potential
(C) Dielectric constant
D. Electrical conductance

Q-3. Which of the following technique is based on the emission of light radiation?
$X$
A. Flame photometry
B. Atomic absorption spectrophotometry
C. Ramen spectroscopy
D. Conductometry

Q-4. Which of the following method is based on the solubility difference between the analyte and the unwanted components

A. Distillation
B. Complex formation
C. Electrodeposition
D. Precipitation

Q-5. Which of the following technique is based on the rate of reaction?

A. Isotopic dilution analysis
B. Flow injection analysis
(C) Mass spectrometry
D. None

Q-6. The term accuracy refers to how near the observed value is to
A. Mean value
B. Significant value
(C) Actual value
D. None
Q.7. When HCl is titrated against AaOH the -pH at equivalence point will be
A. None
(B) equal to 7
C. less than 7
D. greater than 7

Q-8. When CH 3 COOH is titrated against NaOH the pH at the end point is
A. None
B. equal to 7
C. less than 7
(D) greater than 7

Q-9. Which of the following is best indicator for titration of $\mathrm{NH}_{4} \mathrm{OH}$ with HCl ?
$\int$ A. Methyl red
(B) Methyl orange
C. Phenolphthalein
D. None

Q-10. Soap is a salt of compound known as
A. Acetic acid
B. Formic acid
C. Fatty acid
D. Amino acid

Q-11. The colour of water in a lake is due to
(A) Excessive growth of sea weeds
B. Algae
C. Grass
D. Other pollution

Q-12. Cement s a mixture of
A Clay and clinker
(B) Clay, limestone and gypsum
C. Limestone and gypsum
D. Limestone and clay

Q-13. The percentage of nitrogen in urea is
A. $36 \%$
B. $46 \%$
(C) $56 \%$
D. $66 \%$

Q-14. The percentage of nitrogen in ammonia is
A. $32 \%$

V $\begin{array}{r}\text { B. } 42 \% \\ \text { C. } 72 \% \\ \text { D. } 82 \%\end{array}$

Q-15. Copper occur in nature as
A. Native
(B) Combined
C. Both
D. None

Q16. The formula of Cryolite is
(A) $\mathrm{Na}_{3} \mathrm{AlF}_{3}$
B. $\mathrm{Na}_{3} \mathrm{AlF}_{4}$
C. $\mathrm{Na}_{3} \mathrm{AlF}_{8}$
D. None

Q-17. Cells with nuclear envelop belongs to
A. Prokaryotes
(B) Eukaryotes
C. Both
D. None

Q-18. The diameter of a typical plant or animal cell is
A. 1 to $5 \mu \mathrm{~m}$
(B) 5 to $10 \mu \mathrm{~m}$
C. 10 to $50 \mu \mathrm{~m}$
D. 5 to $100 \mu \mathrm{~m}$

Q-19. Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by
(A) Davy
B. Liebig
C. Boyle
D. Rouelle

Q-20. Which of the following statement is not true regarding buffers?
A. They are made up of weak acid and its salts (B. They are made up of weak base and its salts C. they tend to resist a change in pH value D. they are made up of strong acid and weak base

Q-21. Relative order of acidity of $\mathrm{HF}, \mathrm{HCl}, \mathrm{HBr}$ and HI acids is
(A) $\mathrm{HCl}|>\mathrm{HBr}>\mathrm{H}|>\mathrm{HF}$
B. $\mathrm{HF}>\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}$
C. $\mathrm{HI}>\mathrm{HB} />\mathrm{HCl}>\mathrm{HF}$
D. HF H $\mathrm{HI}>\mathrm{HCl}>\mathrm{HBr}$

Q-22. The pH of 0.1 M HCl solution is
$\begin{array}{llll}\text { A. } 0.01 & \text { B. } 0.1 & \text { C. } 0.2 & \text { (D. } 1\end{array}$

Q-23. Which of the following amino acid have two carboxylic groups?
A. Aspartic acid
B. Histidine
C. Gluatmine
D. none

Q-24. How many isomeric aldoses are possible for the molecular formula $\mathrm{C}_{6} \mathrm{H}_{12} \mathrm{O}_{6}$
(A) 2
B. 4
C. 6
D. 16

Q-25. Epimers are compounds that differ in
A. functional group
(B) configuration at alpha carbon
C. configuration at any carbon
D. none

Q-26. Select a basic amino acid
A. Glycine
B. Cystine
C. Alanine
(D) Lysine

Q-27. Apoenzyme is
A. Hydrolytic enzyme
B. Oxidative enzyme
C. Coenzyme
(D) Protein part of enzyme after removal of coenzyme

Q-28. A sugar present in DNA is
A. D-ribose
B. D-glucose
C. 2-Deoxy-D-Ribose
D. None

Q-29. Which vitamin is responsible for vision?
A. Retinol


Q-30. Which of the following is periodic propery? C
A. Atomic volume
B. Metallic character
C. Ionization energy
D. All

Q-31. Which of the following has largest size ?
A. $\mathrm{Na}^{+}$
B. Cl
C. $F^{-}$
(D) $\mathrm{Fe}^{2+}$

Q-32. Which group contains elements that exist as monoatomic molecules?
A. 1
B. 2
C. 14
(D.) 18

Q-33. Which of the following element is more metalize?
A. P
B. As
C. Sb
(D) Bi

Q-34. Which of the following factor effect ionization energy?
A. Atomic radius
C. Shiedding effect
B. Nuclear charge
(D) $A$

Q-35. Which of the following element is metalloid?
A. Pb
B. carbon
C. As
(D) Mg

Q-36. Which of the following compound does not following octet rule?
(A) $\mathrm{CS}_{2}$
B. $\mathrm{PBr}_{3}$
C. IBr
D. $\mathrm{BrF}_{5}$
Q.37. Which of the following halides have the lowest melting point?
A. NaCl
B. NaF
C. NaBr
(D) Nal

Q-38. Which of the following halides has zero dipole moment?
A. $\mathrm{NH}_{3}$
B. $\mathrm{Cl}_{2}$
C. $\mathrm{NF}_{3}$
(D) $\mathrm{CHCl}_{3}$

Q-39. The state of hybridization of carbon in CO 2 is?
A. sp
(B) sp 2
C. sp 3
D. $d s p 2$

Q-40. Which pe of the following is non-polar?
A. $\mathrm{CH}_{2} \mathrm{Cl}$
(B) $\mathrm{CCl}_{4}$
C. $\mathrm{CHCl}_{3}$
D. $\mathrm{CHCl}_{3}$

Q-41. Silvite is an ore of..?
A. Ca
(B) Mg
C. Ba
D. K

X
Q-42. Which elements are non metals?
$N$ \& P
B. As \& Sb
C. $\mathrm{Sb} \& \mathrm{Bi}$
D. $\mathrm{Ba} \& \mathrm{Bi}$

Q-43. Among oxides of nitrogen all are gases except
A. N 2 O 5
B. N 2 O
C. NO D. N2O3
$X$
Q-44. All halogens form oxyacid except
A. Flouring $B$. Chlorine $C$. Bromine (D. Iodine

Q-45. Compounds HCN and HNC are
A. Tautomers
C. functional isomers
(D.) Conformers

## Screening test for the appointment of TGT (Contract)

Subject: Chemistry
Note: Choose the correct answer and all questions are compulsory.

Q-1. The process of identifying the component present in a sample is called

(A )Quantitative analysis
B. Qualitative analysis
C. Volumetric analysis
D. Gravimetric analysis

Q-2. Conductometry is based onA. Electric current
B. Electrical potential
C) Dielectric constant
D. Electrical conductance

Q-3. Which of the following technique is based on the emission of light radiation?
A. Flame photometry

Atomic absorption spectrophotometry
Ramen spectroscopy
D. Conductometry

Q-4. Which of the following method is based on the solubility difference between the analyte and the unwanted components

A. Distillation
B. Complex formation
(C) Electrodeposition
D. Precipitation

Q-5. Which of the following technique is based on the rate of reaction?


A. Isotopic dilution analysis
B. Flow injection analysis
C. Mass spectrometry

None
Q-6. The termaccuracy refers to how near the observed value is to

$\frac{28}{\frac{2}{56}+4+12+16}$
$\frac{2800}{32}$
$\frac{32}{256}$
 at equivalence point will be
A. None
B. equal to 7
C. less than 7
D. greater than 7

Q-8. When CH 3 COOH is titrated against NaOH the pH at the end point is
A. None
B. equal to 7
C. less than 7
(0) greater than 7


Q-9. Which of the following is best indicator for titration of $\mathrm{NH}_{4} \mathrm{OH}$ with HCl ?


Q-10. Soap is a salt of compound known as
A. Acetic acid
B. Formic acid
(C) Fatty acid
D. Amino acid


Q-11. The colour of water in a lake is due to
(A) Excessive growth of sea weeds
B. Algae
C. Grass
D. Other pollution


Q-12. Cement s a mixture of
A Clay and clinker
B. Clay, limestone and gypsum
C. Limestone and gypsum
D. Limestone and clay

Q-13. The percentage of nitrogen in urea is


Q-14. The percentage of nitrogen in ammonia is
A. $32 \%$
B. $42 \%$
C. $72 \%$

$1 \mid$ Page
$\mathrm{OH}-\mathrm{C}^{\mathrm{O}} \mathrm{C}-\mathrm{NH}_{2}$

Q-15. Copper occur in nature as
A. Native
B. Combined
C. Both
D. None


Q-16. The formula of Cryolite is
(A) $\mathrm{Na}_{3} \mathrm{AlF}_{3}$
B. $\mathrm{Na}_{3} \mathrm{AlF}_{4}$
C. $\mathrm{Na}_{3} \mathrm{AlF}_{5}$
D. None

Q-17. Cells with nuclear envelop belongs to


Q-18. The diameter of a typical plant or animal cell is
A. 1 to $5 \mu \mathrm{~m}$
(B) 5 to $10 \mu \mathrm{~m}$ C. 10 to $50 \mu \mathrm{~m}$
D. 5 to $100 \mu \mathrm{~m}$


Q-19. Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by


Q-20. Which of the following statement is not true regarding buffers?
A. They are made up of weak cid and its salts B. They are made up of weak base and its salts
C. they tend to resist a change in pH value D. they are made up of strong acid and weak base

Q-21. Relative order of acidity of $\mathrm{HF}, \mathrm{HCl}, \mathrm{HBr}$ and HI acids is
A.) $\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}>\mathrm{HF}$ B. $\mathrm{HF}>\mathrm{HCl}>\mathrm{HBr}>\mathrm{H}$
C. $\mathrm{Hl}>\mathrm{HBr}>\mathrm{HCl}>\mathrm{HF}$

D. $\mathrm{HF}>\mathrm{HI}>\mathrm{HCl}>\mathrm{HBr}$

Q-22. The pH of 0.1 M HCl solution is
A. 0.01
B. 0.1
C. 0.2


Q-23. Which of the following amino acid have two carboxylic groups?
A. Aspartic acid
B. Histidine
C. Gluatmine


Q-24. How many isomeric aldoses are possible for the molecular formula $\mathrm{C}_{6} \mathrm{H}_{12} \mathrm{O}_{6}$
(A.) 2
$B$.
C. 6
D. 16


Q-25. Epimers are compounds that differ in
A. Aunctional group
B. configuration at alpha carbon
C. configuration at any carbon
D. none

Q-26. Select a basic amino acid


Q-27. Apoenzyme is
A. Hydrolytic enzyme
B. Oxidative enzyme
C. Coenzyme

D. Protein part of enzyme after removal of coenzyme

Q-28. A sugar present in DNA is
A. D-ribose
B. D-glucose
C) 2-Deoxy-D-Ribose
D. None

Q-29. Which vitamin is responsible for vision?
A. Retinol
B. Vit D
C. Tocopherols
D. Vit. K

Q-30. Which of the following is periodic propery?

Q-31. Which of the following has large size ?
A. $\mathrm{Na}^{+}$
B. $\mathrm{Cl}^{-}$
C. $F^{-}$
(D. He ${ }^{2+}$

Q-32. Which group contains elements that exist as monoatomic molecules?
A. 1
B. 2
C. 14


Q-33. Which of the following element is more metallic?
A. P
B. As
C. Sb


Q-34. Which of the following factor effect ionization energy?
A. Atomic radius
C. Shedding effect


Q-35. Which of the following element is metalloid?
A. Pb
(B.) carbon
C. As
D. Mg
K

Q-36. Which of the following compound does not following octet rule?
A. $\mathrm{CS}_{2}$
B. $\mathrm{PBr}_{3}$
C. IBr
(D.) $\mathrm{BrF}_{5}$
1

A
Q-46. Cetane is
A. n-hexane
C. n-octane


Q-47. Which among the following is not nucleophile?
A. $\mathrm{CH}_{3} \mathrm{NH}_{3}$
(C. $\mathrm{CH}^{+}$
B. $\mathrm{CH}_{2} \mathrm{CH} 2$

Q-48. Which of the following is strong acid?
A. Benzoic acid
B. m-nitrobenzoic acid
C. o-nitrobenzoic acid
D. p-nitrobenzoic acid


Q-49. Which of the following orbital does not make sense?
A. $6 f$
B. $4 f$
C. 7 s
(D)

Q-50. The maxium value of order of reaction may be?
A. 1
B. 2

D. 4

Q-37. Which of the following halides have the lowest melting point?
A. NaCl
B. NaF
C. NaBr


Q-38. Which of the following halides has zero dipole moment?
A. $\mathrm{NH}_{3}$
(B.) $\mathrm{Cl}_{2}$
C. $\mathrm{NF}_{3}$
D. $\mathrm{CHCl}_{3}$


Q-39. The state of hybridization of carbon in
CO 2 is?
A. $s p$
B. sp 2
(C.) sp 3
D. dsp 2
$>$

Q-40. Which one of the following is non-polar?
A. $\mathrm{CH}_{2} \mathrm{Cl}$
$\mathrm{CCl}_{4}$
C. $\mathrm{CHCl}_{3}$
D. $\mathrm{CHCl}_{3}$

A. Ca
(B.) Mg
C. Ba
D. $K$


Q-42. Which elements are non metals?
A. N
$N \& P \quad B . A s \& S b$
$\begin{array}{lll}\text { C. } \mathrm{Sb} \& \mathrm{Bi} & \mathrm{D} . \mathrm{Ba} \& \mathrm{Bi}\end{array}$

Q-43. Among oxides of nitrogen all are gases
except
A. N 2 O 5
B. N 2 O
C. NO
(D) ${ }^{2} 2 \mathrm{O} 3$


Q-44. All halogens form oxyacid except
A. Flouring B. Chlorine C. Bromine (D. iodine


Q-45. Compounds HCN and HNC are
A. Tautomer
C. functional isomers

$3 \mid p a g e$

$$
\begin{aligned}
& \text { Name } \\
& \text { Father Nan de }
\end{aligned}
$$

Sunctar kelnmequ

## Screening test for the appointment of TGT (Contract)

 Subject: ChemistryNote: Choose the correct answer and all questions are compulsory.

Q-1. The process of identifying the component present in a sample is called

Quantitative analysis
(B) Qualitative analysis
C. Volumetric analysis
D. Gravimetric analysis

Q-2. Conductometry is based on
A. Electric eurrent
B. Electrical potential
C. Dielectric constant
(D) Electrical conductance

Q-3. Which of the following technique is based on the emission of light radiation?
(A) Flame photometry
B. Atomic absorption spectrophotometry
C. Raman spectroscopy
D. Conductometry

Q-4. Which of the following method is based on the solubility difference between the analyte and the unwanted components
(A.) Distillation
B. Complex formation
C. Electrodeposition
D. Precipitation

Q-5. Which of the following technique is based on the rate of reaction?
A. Isotopic dilution analysis
B. Flow injection analysis
(C. Mass spectrometry
D. None

Q-6. The term accuracy refers to how near the observed value is to
A. Mean value
B. Significant value
C) Actual value
D. None

Q-7. When HCl is titrated against NaOH the pH at equivalence point will be,
A. None
B. equal to 7
C. less than 7
(D) greater than 7

Q-8. When CH 3 COOH is titrated against NaOH the pH at the end point is
(A.) None
B. equal to 7
C. less than 7
D. greater than 7

Q-9. Which of the following is best indicator for titration of $\mathrm{NH}_{4} \mathrm{OH}$ with HCl ?
A. Methyl red
(B) Methyl orange
C. Phenolphthaleine
D. None

Q-10. Soap is a salt of compound known as
(A.) Acetic acid
B. Formic acid
C. Fatty acid

D. Amino acid

Q-11. The colour of water in a lake is due to
A. Excessive growth of sea weeds
B. Algae
C. Grass
D.) Other pollution


Q-12. Cement sa mixture of
(A) Clay and clinker
B. Clay, limestone and gypsum
C. Limestone and gypsum
D. Limestone and clay

Q-13. The percentage of nitrogen in urea is
A. $36 \%$
(B. $46 \%$
C. $56 \%$
D. $66 \%$


Q-14. The percentage of nitrogen in ammonia is
(A. $32 \%$
B. $42 \%$
C. $72 \%$
D. $82 \%$


1) age

Q－15．Copper occur in nature as
A．Native
B．Combined
C．Both
D．None


Q－16．The formula of Cryolite is
A． $\mathrm{Na}_{3} \mathrm{AlF}_{3}$
C． $\mathrm{Na}_{3} \mathrm{AlF}_{5}$

D．None

Q－17．Cells with nuclear envelop belongs to
A．Prokaryotes
（B）Eukaryotes
C．Both
D．None
Q－18．The diameter of a typical plant or animal cell is

A． 1 to $5 \mu \mathrm{~m}$
（B． 5 to $10 \mu \mathrm{~m}$
C． 10 to $50 \mu \mathrm{~m}$


D． 5 to $100 \mu \mathrm{~m}$
Q－19．Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by
A．Davy
B．Liebig
C．Boyle
D．Rouelle


Q－20．Which of the following statement is not true regarding buffers？

A．They are made up of weak acid and its salts B．They are made up of weak base and its salts
C They tend to resist a change in pH yalue D．they are made up of strong acid and weak base

Q－21．Relative order of acidity of $\mathrm{HF}, \mathrm{HCl}, \mathrm{HBr}$ and Hl acids is
A． $\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}>\mathrm{HF}$
（B．） $\mathrm{HF}>\mathrm{HCl}>\mathrm{HBr}>\mathrm{HI}$
C． $\mathrm{HI}>\mathrm{HBr}>\mathrm{HCl}>\mathrm{HF}$
D． $\mathrm{HF}>\mathrm{HI}>\mathrm{HCl}>\mathrm{HBr}$


Q－22．The pH of 0.1 M HCl solution is $\begin{array}{ll}\text {（A．} 0.01 & \text { B．} 0.1\end{array}$ C． 0.2


Q－23．Which of the following amino acid have two carboxylic groups？

A．Aspartic acid
B．Histidine
C．Gluatmine
D．none


Q－24．How many isomeric aldoses are possible for the molecular formula $\mathrm{C}_{6} \mathrm{H}_{12} \mathrm{O}_{6}$
A． 2
B． 4
C． 6
（D．） 16


Q－25．Epimers are compounds that differ in
（A．functional group
B．configuration at alpha cafrodn
C．configuration at any carbon
D．none
Q－26．Select a basic amino acid
A．Glycine
B．Cystine
C．Alanine
（D．Lysine


Q－27．Apoenzyme is
A．Hydrolytic enzyme
B．Oxidative enzyme
C．Coenzyme
D．Protein part of enzyme after removal of coenzyme

Q－28．A sugar present in DNA is


Q－29．Which vitamin is responsible for vision？
（A）Retinol
B．Vit D
C．Tocopherols
D．Vit．K
Q－30．Which of the following is periodic propery？
A．Atomic volume
B．Metallic character
C．Ionization energy
D．All

Q-31. Which of the following has largest size?
A. $\mathrm{Na}^{+}$
B. $\mathrm{Cl}^{-}$
C. $F^{-}$
(D.) $\mathrm{Fe}^{2+}$

V
Q-32. Which group contains elements that exist as monoatomic molecules?
A. 1
B. 2
C. 14
(D) 18

Q-46. Cetane is
A. n-hexane
B. n-pentadecane
C. n-octane
(D. n-hexadecane

Q-47. Which among the following is not nucleophile?

Q-33. Which of the following element is more metallic?
A. P
B. As
C. Sb
(D) Bi

Q-34. Which of the following factor effect ionization energy?
A. Atomic radius
B. Nuclear charge
(C) Shiedding effect
D. All


Q-35. Which of the following element is
metalloid?
A. Pb
(B) carbon
C. As
D. Mg


Q-36. Which of the following compound does not following octet rule?
A $\mathrm{CS}_{2}$
B. $\mathrm{PBr}_{3}$
C. IBr
D. $\mathrm{BrF}_{5}$

Q-37. Which of the following halides have the lowest melting_point?
A. NaCl
(B) NaF
C. NaBr
D. NaI


Q-38. Which of the following halides has zero
dipole moment?
(A.) $\mathrm{NH}_{3}$
B. $\mathrm{Cl}_{2}$
C. $\mathrm{NF}_{3}$
D. $\mathrm{CHCl}_{3}$

Q-39. The state of hybridization of carbon in CO 2 is?
A. sp
B. sp 2
C. sp 3
D. dsp 2

Q-40. Which one of the following is non-polar?
A. $\mathrm{CH}_{2} \mathrm{Cl}$
(B. $\mathrm{CCl}_{4}$
C. $\mathrm{CHCl}_{3}$
D. $\mathrm{CHCl}_{3}$

Q-41. Silvite is an ore of..?
A. Ca
(B. Mg
C. Ba
D. $K$
A. $\mathrm{CH}_{3} \mathrm{NH} 3$
B. $\mathrm{CH}_{2} \mathrm{CH}_{2}$

Q-48. Which of the following is strong acid?
A. Benzoic acid
C. O-nitrobenzoic acid
B. m-nitrobenzoic acid
D. p-nitrobenzoic acid

Q-49. Which of the following orbital does not make sense?
A. $6 f$
B. $4 f$
C. 7 s
(D) 5 g


Q-50. The maxium value of order of reaction may be?
A. 1
B. 2
C. 3
D. 4


Q-42. Which elements are non metals?
A. $N \& P$
B. As \& Sb
C. Sb \& $\mathrm{Bi} \quad \mathrm{D} . \mathrm{Ba} \& \mathrm{Bi}$

Q-43. Among oxides of nitrogen all are gases
except
A. W2O5
B. N 2 O
C. NO
D. N 2 O 3

Q-44. All halogens form oxyacid except
(A. Flouring
B. Chlorine
C. Bromine D. Iodine


Q-45. Compounds HCN and HNC are
(A. Tautomers
B. Metamer
C. functional isomers
D. Conformers


[^0]:    A.fautomers B. Metamer
    D. Conformers

