## AWARD LIST <u>SCREENING TEST (WRITTEN)</u> FOR THE POST OF TRAINED GRADUATE TEACHERS (TGT) (COTNRACT) <u>UNIVERSITY MODEL SCHOOL, UNIVERSITY OF PESHAWAR</u>

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n. 192

### Dated.20.01.2023

<u>S#</u>	Name of applicant with parentage	Subject	Marks	
1.	Durr-e-Shahwar Zafar D/O Mr. Zafar Mumtaz	Chemistry	30	6
2.	Farah Jamshed D/O Qazi Jamshedadam Ali	Chemistry		
3.	Halima Bibi D/O Sher Bahadar Khan	Chemistry	29	5.
4.	Huma Iqbal D/O Mr. Muhammad Iqbal	Chemistry	15	are well
5.	Laila Gulfam D/O Mr. Gulfam	Chemistry	17	art and a star
6.	Maheen Rahim D/O Mr. Fazal e Rahim	Chemistry	24	4.8
7.	Saima Momin D/O Mr. Momin Khan	Chemistry	23	] (*
8.	Sana Nabi D/O Mr. Muhammad Nabi	Chemistry	33	6.
9.	Shaista Zeb D/O Mr. Abdul Saboor Khan	Chemistry	21	Cy.
10.	Shumaila Javed D/O Mr. Javed Iqbal	Chemistry		
11.	Sumera Rehman D/O Mr. Sher Rehman	Chemistry	20	4
12.	Wagma D/O Mr. Momin Khan	Chemistry	23	4-
13.	Rozina D/O Mr. Sher Daraz Khan	Chemistry	29	5

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## AWARD LIST <u>SCREENING TEST (WRITTEN)</u> <u>FOR THE POST OF TRAINED GRADUATE TEACHERS (TGT) (COTNRACT)</u> <u>UNIVERSITY PUBLIC SCHOOL, UNIVERSITY OF PESHAWAR</u>

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Dated.20.01.2023

<u>S#</u>	Name of applicant with parentage	Subject	Marks	
1.	Bakhtawara D/O Malik Abdul baser Khan	Chemistry		
2.	Huma Iqbal D/O Mr. Muhammad Iqbal	Chemistry	The same from th	TS 3
3.	Muhammad Naveed S/O Mr. Jehangir Khan	Chemistry	18	3.6
4.	Muhammad Saqib S/O Mr. Muhammad Shahab	Chemistry		
5.	Qaisar Khan S/O Mr. Ayub Khan	Chemistry	35	
6.	Rozina D/O Mr. Sher Daraz Khan	Chemistry		(29) 5-8
7.	Saif Ullah Khan S/O Haji Lal Khan	Chemistry	27	54
8.	Sana Nabi D/O Mr. Muhammad Nabi	Chemistry	1	33) 6 -





84 5	
Q-31. Which of the following has largest size ?	0.40 0.
$(\underline{B}, C) = C, F = D, Fe^{2+1} X$	Q-46. Cetane is
Q-32. Which group contains element up	C n-octano B. n-pentadecano
as monoatomic molecules?	D n-hexadecane
A. 1 B. 2 C. 14 (P) 18 .	Q-47. Which among the follow:
Q-33. Which of the f it	nucleophile ?
metallic?	A. CH3NH3 B. CH2CH2
A.P B.As C.Sh	D. OH
	Q-48. Which of the following
Q-34. Which of the following factor effect	A. Benzoic acid B strong acid?
A. Atomic radius	C. o-nitrobenzoic acid
C. Shiedding effect B. Nuclear charge	
BAIL	make sense?
Q-35. Which of the following element in	A. 6f B Af
A Ph	D. 41 C.7s D5g
B. carbon (C.)As D. Mg I	Q-50. The maxium value of order of a
Q-36. Which of the following	A 1 A A A A A A A A A A A A A A A A A A
following octet rule?	$B_2$ (2) 3 D.4 V
A. CS <sub>2</sub> B. PBr <sub>3</sub> C. IBr D. Prr	
Q-37 Which a CH	
lowest melting pointo	*****************
A. NaCI B NaE C NaD	
D Nal	
Q-38. Which of the following halides has zero	
A NH-	
$Cl_2$ C. NF <sub>3</sub> D. CHCl <sub>3</sub>	
Q-39. The state of hybridization	
CO2 is?	
(A) sp B. sp2 C. sp3 D dsp2 A	
<b>Q-40.</b> Which one of the following is non-polar? A. $CH_2CI$ (B. $CCI_4$ C. $CHCI_3$ D. $CHCI_5$ )	
Q-41. Silvite is an one of a	
A. Ca B. Mg C. Ba Dr N	
A N & P B. As & Sb C. Sb & Bi D Bo & Bi	
Q-43. Among oxides of nitrogen all are -	
A NOCE D Mar and a gases	
U 11200 B. N2O C. NO D. N2O3 N	
Q-44. All halogens form	
A. Flourine B. Chlorine C. Bromine Ditedian V	
Q-45. Compunds HCN and HNC are A. Tautomers C. functional isomers D. Conformers	
	3   Page

Name -> Laila Gulfam F. Name -> Gulfam Q-7. When HCI is titrated against NaOH the pH Screening test for the appointment of at equivalence point will be **TGT (Contract)** A. None Subject: Chemistry B. equal to 7 Note: Choose the correct answer and all C. less than 7 questions are compulsory. VD. greater than 7 Q-1. The process of identifying the component Q-8. When CH3COOH is titrated against NaOH present in a sample is called the pH at the end point is VA. Quantitative analysis A. None B. Qualitative analysis B. equal to 7 C. Volumetric analysis Ve. Tess than 7 D. Gravimetric analysis D. greater than 7 Q-2. Conductometry is based on Q-9. Which of the following is best indicator for titration of NH₄OH with HCI? A. Electric current B. Electrical potential A. Methyl red C. Dielectric constant **VB**. Methyl orange D. Electrical conductance C. Phenolphthaleine D. None Q-3. Which of the following technique is based on the emission of light radiation? Q-10. Soap is a salt of compound known as A. Flame photometry A. Acetic acid VB. Atomic absorption spectrophotometry B. Formic acid C. Raman spectroscopy VC. Fatty acid D. Conductometry D. Amino acid Q-4. Which of the following method is based on Q-11. The colour of water in a lake is due to the solubility difference between the analyte and the unwanted components A. Excessive growth of sea weeds B. Algae A. Distillation C. Grass B. Complex formation D. Other pollution C. Electrodeposition **D**. Precipitation Q-12. Cement s a mixture of Q-5. Which of the following technique is based A. Clay and clinker on the rate of reaction? B. Clay, limestone and gypsum C. Limestone and gypsum A. Isotopic dilution analysis D. Limestone and clay B. Flow injection analysis C. Mass spectrometry Q-13. The percentage of nitrogen in urea is D. None A. 36% Q-6. The term accuracy refers to how near the **B.** 46% observed value is to C. 56% D. 66% A. Mean value B. Significant value Q-14. The percentage of nitrogen in ammonia is ✓C. Actual value A. 32% D. None B. 42% Ve. 72% D. 82% 1 | Page



2 Page

Q-31. Which of the following has largest size ? A. Na<sup>+</sup> B. Cl VC/F D. Fe<sup>2+</sup> Q-32. Which group contains elements that exist as monoatomic molecules? A.1 B. 2 C. 14 D. 18 Q-33. Which of the following element is more metallic? A. P C.Sb B. As D. Bi Q-34. Which of the following factor effect ionization energy? A. Atomic radius B. Nuclear charge C. Shiedding effect D. All 4-35. Which of the following element is metalloid? A. Pb B. carbon C. As D. Ma Q-36. Which of the following compound does not following octet rule? A. CS<sub>2</sub> B. PBr<sub>3</sub> C IBr D. BrF<sub>5</sub> Q-37. Which of the following halides have the lowest melting point? A. NaCl B. NaF C. NaBr D. Nal Q-38. Which of the following halides has zero dipole moment? B. Cl<sub>2</sub> C. NF<sub>3</sub> A. NH<sub>3</sub> D. CHCl<sub>3</sub> Q-39. The state of hybridization of carbon in CO2 is? B. sp2 A. sp C. sp3 D. dsp2 Q-40. Which one of the following is non-polar? A. CH2CI B. CCI4 C. CHCl<sub>3</sub> D. CHCl<sub>3</sub> Q-41. Silvite is an ore of ..? A. Ca B. Mg C. Ba Ъ. к Q-42. Which elements are non metals? 🗸 Ń&P B. As&Sb C. Sb&Bi D. Ba&Bi Q-43. Among oxides of nitrogen all are gases except A. N2O5 B. N2O C. NO D. N2O3 Q-44. All halogens form oxyacide except A. Flourine B. Chlorine C. Bromine D. Iodine Q-45. Compunds HCN and HNC are A. Tautomers **B.** Metamers C. functional isomers D. Conformers



\*\*\*\*\*

Name Sana Nab Father Name Muhammad Nabi EUMS : S.NO # 2 TGT UPS Screening test for the appointment of Q-7. When HCI is titrated against NaOH the pH at equivalence point will be TGT (Contract) A. None Subject: Chemistry Prequal to 7 Note: Choose the correct answer and all C. less than 7 questions are compulsory. D. greater than 7 Q-1. The process of identifying the component Q-8. When CH3COOH is titrated against NaOH present in a sample is called The pH at the end point is A. Quantitative analysis A. None B. Qualitative analysis B. equal to 7 C. Volumetric analysis C. less than 7 D. Gravimetric analysis DY greater than 7 Q-2. Conductometry is based on Q-9. Which of the following is best indicator for titration of NH₄OH with HCI? A. Electric current B. Electrical potential A. Methyl red C. Dielectric constant B. Methyl orange Blectrical conductance C. Phenolphthaleine D. None Q-3. Which of the following technique is based of the emission of light radiation? Q-19. Soap is a salt of compound known as K. Flame photometry A. Acetic acid B Atomic absorption spectrophotometry B. Formic acid C. Raman spectroscopy Fatty acid D. Conductometry D. Amino acid Q.4. Which of the following method is based on 0.11. The colour of water in a lake is due to The solubility difference between the analyte and the unwanted components A. Excessive growth of sea weeds Algae A. Distillation C. Grass B. Complex formation D. Other pollution C. Electrodeposition D. Precipitation Q-12. Cement s a mixture of Q-5. Which of the following technique is based 🖌 Clay and clinker on the rate of reaction? B. Clay, limestone and gypsum C. Limestone and gypsum A Isotopic dilution analysis D. Limestone and clay **VB**. Flow injection analysis C. Mass spectrometry Q-13. The percentage of nitrogen in urea is D. None A. 36% Q-6. The term accuracy refers to how near the B. 46% observed value is to C. 56% **VET.** 66% A. Mean value B. Significant value Q-#4. The percentage of nitrogen in animonia is C. Actual value Á. 32% D. None B. 42% C.72% **D**. 82%



D. None

Q-16. The formula of Cryolite is

A. Na<sub>3</sub>AIF<sub>3</sub> B. Na<sub>3</sub>AIF<sub>4</sub>

- C. Na<sub>3</sub>AIF<sub>5</sub>
- D. None

Q-17. Cells with nuclear envelop belongs to

A. Prokaryotes 1

✓B. Eukaryotes

C. Both

D. None

Q-18. The diameter of a typical plant or animal cell is

A. 1 to 5 µm **√**8. 5 to 10 µm C. 10 to 50 µm D. 5 to 100 µm /

Q-19. Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by

🖌 Davv B. Liebig C. Boyle D. Rouelle

Q-20. Which of the following statement is not true regarding buffers?

A. They are made up of weak acid and its salts B. They are made up of weak base and its salts C. they tend to resist a change in pH value D they are made up of strong acid and weak base

Q-21. Relative order of acidity of HF, HCI, HBr and HI acids is

A. HCI>HBr>HI>HF

B. HF>HCI>HBr>HI

Ve. HI>HBr>HCI>₩

D. HF>HI>HGt>HBr

Q-22-The pH of 0.1 M HCl solution is A. 0.01 B. 0.1 C. 0.2 D. 1

Q-23. Which of the following amino acid have two carboxylic groups?

X. Aspartic acid

- B. Histidine
- C. Gluatmine
- D. none

Q-24. How many isomeric aldoses are possible for the molecular formula C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>



Q-25. Epimers are compounds that differ in

A. functional group

- ✓B configuration at alpha carbon
- C. configuration at any carbon
- D. none

Q-26. Select a basic amino acid



Q-27. Apoenzyme is

- A. Hydrolytic enzyme
- B. Oxidative enzyme
- Coenzyme

D. Protein part of enzyme after removal of coenzyme

Q-28. A sugar present in DNA is

- A. D-ribose
- B. D-glucose
- √C. 2-Deoxy-D-Ribose
- D. None

Q-29. Which vitamin is responsible for vision?



- A. Atomic volume C. Ionization energy VD. All
- 2 | Page

Q-31. Which of the following has largest size ? Q-46. Cetane is A. Na<sup>+</sup> B. Cl VC. F D. Fe<sup>24</sup> A. n-hexane B. n-pentadecane C. n-octane D. n-hexadecane Q-32. Which group contains elements that exist as monoatomic molecules? Q-47. Which among the following is not V. 18 A. 1 B. 2 C. 14 fucleophile ? A. CH3NH3 B. CH2CH2 Q-33. Which of the following element is more √Q. CH3+ D. OH metallic? √б. Ві A.P B. As C.Sb Q-48. Which of the following is strong acid? A. Benzoic acid ∼B. m-nitrobenzoic acid Q-34. Which of the following factor effect C. o-nitrobenzoic acid D. p-nitrobenzoic acid ionization energy? A. Atomic radius B. Nuclear charge Q-49. Which of the following orbital does not C. Shiedding effect D. All make sense? A. 6f B. 4f C.7s Q-25. Which of the following element is metalloid? Q-50. The maxium value of order of reaction A. Pb B. carbon V. As D. Ma may be? A. 1 VC. 3 B. 2 Q-36. Which of the following compound does not following octet rule? A. CS<sub>2</sub> B. PBr<sub>3</sub> C. IBr VØ, BrF₅ Q-37. Which of the following halides have the Kwest melting point? A. NaCL B. NaF C. NaBr VD. Nal Q-38. Which of the following halides has zero dipole moment? A. NH<sub>3</sub> VB. Ch C. NF<sub>3</sub> D. CHCl<sub>3</sub> Q-39. The state of hybridization of carbon in CO2 is? A. sp B. sp2 C. sp3 D. dsp2 Q40. Which one of the following is non-polar? A. CH<sub>2</sub>CL VS. CCI4 C. CHCI<sub>3</sub> D. CHCla Q-41. Silvite is an ore of ..? A. Ca J. Mg C. Ba D. K Q-42. Which elements are non metals? A.N&P B.As&Sb C. Sb & Bi D. Ba & Bi Q-43. Among oxides of nitrogen all are gases except XX. N2O5 B. N2O C, NO D. N2O3 Q-44. All halogens form oxyacide except A. Flourine B. Chlorine C. Bromine V. Iodine Q-45. Compunds HCN and HNC are A. Tautomers B. Metamers C\_functional isomers D. Conformers

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D. Lysine

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- B. Oxidative enzyme
- C. Coenzyme
- (D) Protein part of enzyme after removal of coenzyme

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	<b>3</b>   P a g e

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Huma labal Nohammad labal

# Screening test for the appointment of TGT (Contract) Subject: Chemistry

**Note:** Choose the correct answer and all questions are compulsory.

**Q-1.** The process of identifying the component present in a sample is called

A. Quantitative analysis

- B. Qualitative analysis
- VerVolumetric analysis
- D. Gravimetric analysis

Q-2. Conductometry is based on

- A. Electric current
- B. Electrical potential
- C. Dielectric constant

LD Electrical conductance

**Q-3.** Which of the following technique is based on the emission of light radiation?

A. Flame photometry

- B Atomic absorption spectrophotometry
- C. Raman spectroscopy
- D. Conductometry

**Q-4.** Which of the following method is based on the solubility difference between the analyte and the unwanted components

- A. Distillation B. Complex formation C. Electrodeposition
  - D. Precipitation

**Q-5.** Which of the following technique is based on the rate of reaction?

- A. Isotopic dilution analysis B. Flow injection analysis
- -C. Mass spectrometry
- D. None

Q-6. The term accuracy refers to how near the observed value is to



- B. Significant value
- C. Actual value
  - D. None

- Q-7. When HCl is titrated against NaOH the pH
- at equivalence point will be A. None
- B\_equal to 7
- C less than 7
- D. greater than 7

Q-8. When CH3COOH is titrated against NaOH the pH at the end point is

Û

- A. None
- B. equal to 7
- C. less than 7
- D: greater than 7

**Q-9.** Which of the following is best indicator for titration of  $NH_4OH$  with HCI?

- A. Methyl red
- B. Methyl orange
- Le Phenolphthaleine



Q-10. Soap is a salt of compound known as

- A. Acetic acid
- LB. Formic acid
- C. Fatty acid
- D. Amino acid
- Q-11. The colour of water in a lake is due to
- A. Excessive growth of sea weeds
- B. Algae
  - C. Grass
  - D. Other pollution

Q-12. Cement s a mixture of

- A. Clay and clinker
- B. Çlay, limestone and gypsum
- LC. Limestone and gypsum
  - D. Limestone and clay

Q-13. The percentage of nitrogen in urea is

A. 36% V B. 46% C. 56%

D. 66%

Q-14. The percentage of nitrogen in ammonia is





2 Page

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- Q-15. Copper occur in nature as A. Native (B) Combined
- C Both D. None

Q-16. The formula of Cryolite is



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- D nono
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- A. functional group
- B. configuration at alpha carbon
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A. Atomic volume B. Metallic character

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C. Ionization energy D All

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κ.				7-201-	085	387-	6.				Ŧ
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<b>Q-3</b> as r A. 1	<b>32.</b> Which ( monoatom 1	group con iic molecu B. 2	tains elen les? C. 14	nents that e	exist	Q-47. White nucleophile	ch among tl e ?	he followin	g is not		
Q-3	33. Which (	of the follo	wing elen	nent is mor	e	A. CH3NH	3	B. CH2 D. OH	CH2		
A. F	D	B. As	C.Sb	Ві	$\bigvee$	<b>Q-48.</b> Whic A. Benzoic	ch of the fol acid	lowing is s	trong acid	l? n acid	X
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A. A C. S	Atomic rad Shiedding	ius effect	B. Nucl	lear charge		Q-49. Whice make sense	ch of the fol e? B 4f	lowing orb	ital does r	not	and the second
Q-3 met	5. Whi <mark>ch c</mark> talloid?	of the follo	wing elen	nent is		<b>Q-50</b> . The	maxium val	U.7S	r of reaction	on	
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follo A. C	wing octel	t rule? B. PBr₃	wing com C. IBr	D BrF₅ \	s not						
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<b>Q-3</b> 9 CO2	9. The stat 2 is?	te of hybri	dization o	f carbon in		Management -					
(A)s;	p B.	sp2 C	sp3	D. dsp2		and the second					
<b>Q-4</b> ( A C	0. Which o CH₂CI (B)	ne of the t CCl₄ (	following i C. CHCl <sub>3</sub>	is non-pola D. CHC	17? 213 L			- 5- V			
<b>Q-4</b> 1 A. C	1. Silvite is a B	an ore of ≽Mg C.	? Ba D.	KV			L. H.				
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SAIF Ullah

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MAHEEN RAHIM FAZAC-NE-AHINA 17301-9730205-4 Q-7. When HCI is titrated at ainst NaOH the pH Screening test for the appointment of at equivalence point will be TGT (Contract) A. None Subject: Chemistry B. equal to 7 Note: Choose the correct answer and all C. less than 7 questions are compulsory. D) greater than 7 C Q-1. The process of identifying the component Q-8. When CH3COOH is titrated against NaOH present in a sample is called the pH at the end point is A<sup>J</sup>Quantitative analysis A. None B. Qualitative analysis B. equal to 7 C. Volumetric analysis C)less than 7 D. Gravimetric analysis D. greater than 7 Q-2. Conductometry is based on Q-9. Which of the following is best indicator for titration of NH4OH with HCI? A. Electric current B. Electrical potential Methyl red C. Dielectric constant B)Methyl orange DElectrical conductance C. Phenolphthaleine D. None Q-3. Which of the following technique is based on the emission of light radiation? Q-10. Soap is a salt of compound known as A. Flame photometry A. Acetic acid (B) Atomic absorption spect/ophotometry B. Formic acid C. Raman spectroscopy (C) Fatty acid D. Conductometry D. Amino acid Q-4. Which of the following method is based on Q-11. The colour of water in a lake is due to the solubility difference between the analyte and the unwanted components (A) Excessive growth of sea weeds B. Algae A. Distillation C. Grass B. Complex formation D. Other pollution C. Electrodeposition (D)Precipitation Q-12. Cement s a mixture of Q-5. Which of the following technique is based A, Clay and clinker on the rate of reaction? B. Clay, limestone and gypsum C. Limestone and gypsum A. Isotopic dilution analysis D. Limestone and clay B. Flow injection analysis C. Mass spectrometry Q-13. The percentage of nitrogen in urea is D. None 36% Q-6. The term accuracy refers to how near the B. 46% observed value is to C. 56% D. 66% A. Mean value B Significant value Q-14. The percentage of nitrogen in ammonia is C. Actual value A. 32% D. None B. 42% <u>Ç</u>. 72%

(D.)82%

<sup>1 |</sup> Page



Q-16. The formula of Cryolite is



Q-17. Cells with nuclear envelop belongs to

A. Prokaryotes B.Eukaryotes C. Both D. None

**Q-18.** The diameter of a typical plant or animal cell is

A. 1 to 5 μm B. 5 to 10 μm C. 10 to 50 μm D 5 to 100 μm



**Q-19.** Acids are substances whose aqueous solutions turned blue litmus red and tasted sour stated by

A Davy B. Liebig

- C. Boyle
- D. Rouelle

**Q-20**. Which of the following statement is not true regarding buffers?

A. They are made up of weak acid and its salts B. They are made up of weak base and its salts C. they tend to resist a change in pH value D; they are made up of strong acid and weak base

Q-21. Relative order of acidity of HF, HCI, HBr and HI acids is A. HCI>HBr>HI>HF

pll = - of fuil

B. HF>HCI>HBr>HI C HI>HBr>HCI>HF

D. HF>HI>HCI>HBr

**Q-22.** The pH of 0.1 M HCl solution is A. 0.01 B. 0.1 C. 0.2 (D.)1

**Q-23.** Which of the following amino acid have two carboxylic groups?

Aspartic acid B. Histidine C. Gluatmine

D. none

Q-24. How many isomeric aldoses are possible for the molecular formula  $C_6H_{12}O_6$ 

(A) 2 B. 4 C. 6 D. 16

Q-25. Epimers are compounds that differ in

A. functional group

- B. configuration at alpha carbon
- C configuration at any carbon
- D. none

Q-26. Select a basic amino acid

- A. Glycine B. Cystine <u>C</u>. Alanine
- D. Lysine

Q-27. Apoenzyme is

- A. Hydrolytic enzyme
- B. Oxidative enzyme
- C. Coenzyme

D Protein part of enzyme after removal of coenzyme

Q-28. A sugar present in DNA is

- A. D-ribose
- B. D-glucose (C.)2-Deoxy-D-Ribose
- D. None

Q-29. Which vitamin is responsible for vision?



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Q-31. Which A. Na'	of the foll B. Cl <sup>°</sup>	owing has I C. F⁻	argest si: D. Fe <sup>2+</sup>	ze ?	<b>Q-46.</b> Cet An-hexa C. n-octa	ane is ne ne	B. n-pe D. n-he	ntadec: xadeca	ane X
<b>Q-32.</b> Which as monoaton A, 1	group cor nic molect B. 2	ntains elem ules? (C) 14	ents that D. <b>18</b>	exist	<b>Q-47.</b> Whi nucleophil	ich among th	e followin	g is not	
<b>Q-33.</b> Which metallic?	of the foll	owing elem	ent is mo	ore	A. CH3NH C. CH3+	13	B. CH2 D. OH	CH2	
A. P <b>Q-34.</b> Which	B. As	C.Sb owing facto	D Bi		Q-48. Whi A. Benzoid	ch of the follo c acid	owing is s B. m-ni	trong a trobenz	cid? oic acid
Ionization ene A. Atomic rac	ergy? dius	B. Nucle	ear charg	e	Q-49. Whi	ch of the follo	D. p-niti wing orbi	robenzo ital doe	s not
Q-35. Which	of the folk	owing elem	∿ ent is		make sens A. 6f	se? B. 4f	C.7s	Dбg	
Metalloid?	B. carbon	C. As	D. Mg	X	<b>Q-50.</b> The may be?	maxium valu	e of orde	r of rea	ction
Q-36. Which following octe	of the folk et rule? B_PBre	wing comp	ound doe	es not	/ \. 1	D. Z	(C.)3	U. 4	
Q-37. Which a	of the folic	owing halide	es have th	ne	*********	******	******	******	****
A. NaCl	B. NaF	C. NaBr	D Nal						
dipole momer A. $NH_3$ (	nt? B_Cl₂	C. NF <sub>3</sub>	D. CHCla	s X				<i>u</i>	Sp 1
<b>Q-39.</b> The sta CO2 is? A. sp B.	ite of hybr . sp2	idization of	carbon ir ). dsp2	$\times$		0 = 1		2	ţ
<b>Q-40.</b> Which $c$ A. CH <sub>2</sub> Cl $(B)$	one of the CCl₄	following is C. CHCl₃	non-pola D. CH	ar? Cl <sub>3</sub>					
<b>Q-41.</b> Silvite is A. Ca	s an ore o Mg C	f? ). Ba — D. I	к 1						
<b>Q-42.</b> Which e	elements a As & Sb	are non met C. Sb & E	als? 3i D.Ba	& ві 🖌					
Q-43. Among except A. N2O5 B.	oxides of N2O C	nitrogen all	are gase	es	****				
<b>Q-44.</b> All halog A. Flourine B	gens form 5. Chlorine	oxyacide e C. Bromi	except ne D. loc	dine V					
Q-45. Compur A. Tautomers C. functional is	nds HCN a somers	and HNC ar B. Meta D. Confor	re mers mers	X					
								3	Page

Wagma Momen F/N Momen Khan Q.7. When HCI is titrated against NaOH the pH Screening test for the appointment of at equivalence point will be TGT (Contract) A. None Bequal to 7 Subject: Chemistry Note: Choose the correct answer and all C. less than 7 questions are compulsory. D. greater than 7 Q-1. The process of identifying the component Q-8. When CH3COOH is titrated against NaOH present in a sample is called the pH at the end point is A Quantitative analysis A. None B. Qualitative analysis B. equal to 7 C. Volumetric analysis C. less than 7 D. Gravimetric analysis (D) greater than 7 Q-2. Conductometry is based on Q-9. Which of the following is best indicator for titration of NH₄OH with HCI? A. Electric current B. Electrical potential A. Methyl red (C) Dielectric constant @Methyl orange D. Electrical conductance C. Phenolphthaleine D. None Q-3. Which of the following technique is based on the emission of light radiation? Q-10. Soap is a salt of compound known as A. Flame photometry A. Agetic acid B. Atomic absorption spectrophotometry B. Formic acid (C)Raman spectroscopy C Fatty acid D. Conductometry D. Amino acid Q-4. Which of the following method is based on Q-11. The colour of water in a lake is due to the solubility difference between the analyte and the unwanted components A Excessive growth of sea weeds B. Algae A. Distillation C. Grass B. Complex formation D. Other pollution (C) Electrodeposition D. Precipitation Q-12. Cement s a mixture of Q-5. Which of the following technique is based , Elay and clinker on the rate of reaction? B) Clay, limestone and gypsum C. Limestone and gypsum A. Isotopic dilution analysis D. Limestone and clay B. Flow injection analysis C)Mass spectrometry Q-13. The percentage of nitrogen in urea is D. None A. 36% Q-6. The term accuracy refers to how near the B. 46% observed value is to **(C)** 56% D. 66% A. Mean value B. Significant value Q-14. The percentage of nitrogen in ammonia is C Actual value A. 32% D. None B. 42% 72% (D.)82% 1 | Page



<sup>2 |</sup> Page

Q-31. Which of the following has largest size ? Q-46. Cetane is A Na<sup>+</sup> B. CI C. F  $(D)Fe^{2+}$ A. n-hexane B. n-pentadecane C. n-octane (D)n-hexadecane Q-32. Which group contains elements that exist as monoatomic molecules? Q-47. Which among the following is not A. 1 B. 2 C. 14 (D)18 nucleophile? A. CH3NH3 B. CH2CH2 Q-33. Which of the following element is more (C)CH3+ D. OH metallie? A.P B. As C.Sb D) Bi Q-48. Which of the following is strong acid? A. Benzoic acid (B) m-nitrobenzoic acid Q-34 Which of the following factor effect C. o-nitrobenzoic acid D. p-nitrobenzoic acid ionization energy? A. Atomic radius B. Nuclear charge Q-49. Which of the following orbital does not C. Shiedding effect (D) All make sense? A. 6f B. 4f (D)5g C.7s Q-35. Which of the following element is metalloid? Q-50. The maxium value of order of reaction A. Pb B. carbon C. As (D)Mg may be? A. 1 B. 2 C) 3 D. 4 Q-36. Which of the following compound does not following octet rule? (A) CS<sub>2</sub> B. PBr<sub>3</sub> C. IBr D. BrF<sub>5</sub> \*\*\*\*\*\*\*\*\*\* Q-37. Which of the following halides have the fowest melting point? A. NaCl B. NaF C. NaBr /d')Nal Q-38. Which of the following halides has zero dipole moment? A. NH<sub>3</sub> B. Cl<sub>2</sub> C. NF<sub>3</sub> (D) CHCI3 Q-39. The state of hybridization of carbon in CO2 is? A. sp (B) sp2 C. sp3 D. dsp2 Ø-40. Which one of the following is non-polar? A.  $CH_2CI$  (B)  $CCI_4$ C. CHCl<sub>3</sub> D. CHCl<sub>3</sub> & 41. Silvite is an ore of ..? A. Ca (B)Mg C Ba D. K Q-42. Which elements are non metals? (A.)N & P B. As & Sb C. Sb & Bi D. Ba & Bi Q-43. Among oxides of nitrogen all are gases except A. N2O5 B. N2O C.NO D.)N2O3 Q-44. All halogens form oxyacide except A. Flourine B. Chlorine C. Bromine(D)lodine Q-45. Compunds HCN and HNC are A. Tautomers B. Metamers C. functional isomers (D.)Conformers





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Q-32. Which group contains elements that exist as monoatomic molecules?	e
A.1 B.2 C.14 DY18 in public Participation antioning the following is not	_
Q-33. Which of the following element is more metallic?	L
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Q-35. Which of the following element is metalloid? A. Pb B. carbon C. As D. Mg A. or B. 4f C. 7s DBg Q-50. The maxium value of order of react may be?	tion
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3   P	age
6 P Ġ	
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### Screening test for the appointment of TGT (Contract) Subject: Chemistry

**Note:** Choose the correct answer and all questions are compulsory.

**Q-1.** The process of identifying the component present in a sample is called

K. Quantitative analysis

. Lin

- B Qualitative analysis
- C. Volumetric analysis
- D. Gravimetric analysis

Q-2. Conductometry is based on

- A. Electric eurrent
- B. Electrical potential

C. Dielectric constant

D Electrical conductance

**Q-3.** Which of the following technique is based on the emission of light radiation?

K Flame photometry

- B. Atomic absorption spectrophotometry
- C. Raman spectroscopy
- D. Conductometry

**Q-4.** Which of the following method is based on the solubility difference between the analyte and the unwanted components

(A) Distillation

- B. Complex formation
- C. Electrodeposition
  - D. Precipitation

**Q-5.** Which of the following technique is based on the rate of reaction?

A. Isotopic dilution analysis

- B. Flow injection analysis
- C. Mass spectrometry

D. None

Q-6. The term accuracy refers to how near the observed value is to

A. Mean value B. Significant value C Actual value D. None

Sumera Rehman

Q-7. When HCl is titrated against NaOH the pH at equivalence point will be

C

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- A. None
- B. equal to 7
- C. less than 7
- D greater than 7

Q-8. When CH3COOH is titlated against NaOH the pH at the end point is

A.)None B. equal to 7 C. less than 7 D. greater than 7

**Q-9.** Which of the following is best indicator for titration of  $NH_4OH$  with HCl?

- A. Methyl red
- B Methyl orange
- C. Phenolphthaleine
- D. None

Q-10. Soap is a salt of compound known as

- (A) Acetic acid
- B. Formic acid
- C. Fatty acid
- D. Amino acid
- Q-11. The colour of water in a lake is due to
- A. Excessive growth of sea weeds
- B. Algae
- C. Grass
- (D.) Other pollution
- Q-12. Cement s a mixture of
- (A) Clay and clinker
- B. Clay, limestone and gypsum/
- C. Limestone and gypsum
- D. Limestone and clay

Q-13. The percentage of nitrogen in urea is

- A. 36% (B) 46%
- C. 56%
- D. 66%

**Q-14.** The percentage of nitrogen in ammonia is



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Q-16. The formula of Cryolite is



Q-17. Cells with nuclear envelop belongs to

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>HBr>H	
>HCI>HF	1
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- B. configuration at alpha carbon
- C. configuration at any carbon \
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- A.)Hydrolytic enzyme
- B. Oxidative enzyme
- C. Coenzyme

D. Protein part of enzyme after removal of coenzyme

Q-28. A sugar present in DNA is



D. None

Q-29. Which vitamin is responsible for vision?



Name

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Sumera Robineco

Q-31. Which of the following has largest size ?	Q-46. Cetane is
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	***************************************
G-57. Which of the following halides have the	
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0 41 Silvita is an are of 2	
$\Delta$ Ca $(\mathbf{R}^{2})$ Ma $(\mathbf{R}^{2})$ $(\mathbf{R}^{2})$	
A. Od (D. Mig C. Da D. K	
<b>Q-42</b> . Which elements are non metals?	<b>,</b>
(A) N & P B As & Sb C Sb & Bi D Ba & Bi X	
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